

13

ELECTRONS IN ATOMS

Conceptual Curriculum

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☐ More abstract concepts or math/problem-solving

Standard Curriculum

- ☒ Core content
☐ Extension topics

Honors Curriculum

- ☒ Core honors content
☐ Options for acceleration

SECTION 13.1 MODELS OF THE ATOM (pages 361–366)

This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

► The Evolution of Atomic Models (pages 361–363)

1. What are the chemical properties of atoms, ions, and molecules related to?

The properties are related to the arrangement of electrons.

2. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

3. The energy level of an electron is the region around the nucleus where the electron is likely to be moving.
4. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false
5. Circle the letter of the term that completes the sentence correctly. A quantum of energy is the amount of energy required to
- move an electron from its present energy level to the next lower one
 - maintain an electron in its present energy level
 - ☒ move an electron from its present energy level to the next higher one
6. In general, the higher the electron is on the energy ladder, the farther it is from the nucleus.

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David Baud



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Table of Contents Prentice Hall Chemistry Workbook Answers Ch 13

1. Understanding the eBook Prentice Hall Chemistry Workbook Answers Ch 13
 - The Rise of Digital Reading Prentice Hall Chemistry Workbook Answers Ch 13
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Workbook Answers Ch 13
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Workbook Answers Ch 13
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Workbook Answers Ch 13
 - Personalized Recommendations
 - Prentice Hall Chemistry Workbook Answers Ch 13 User Reviews and Ratings
 - Prentice Hall Chemistry Workbook Answers Ch 13 and Bestseller Lists
5. Accessing Prentice Hall Chemistry Workbook Answers Ch 13 Free and Paid eBooks
 - Prentice Hall Chemistry Workbook Answers Ch 13 Public Domain eBooks
 - Prentice Hall Chemistry Workbook Answers Ch 13 eBook Subscription Services
 - Prentice Hall Chemistry Workbook Answers Ch 13 Budget-Friendly Options

6. Navigating Prentice Hall Chemistry Workbook Answers Ch 13 eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Workbook Answers Ch 13 Compatibility with Devices
 - Prentice Hall Chemistry Workbook Answers Ch 13 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Workbook Answers Ch 13
 - Highlighting and Note-Taking Prentice Hall Chemistry Workbook Answers Ch 13
 - Interactive Elements Prentice Hall Chemistry Workbook Answers Ch 13
8. Staying Engaged with Prentice Hall Chemistry Workbook Answers Ch 13
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Workbook Answers Ch 13
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Workbook Answers Ch 13
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Workbook Answers Ch 13
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Workbook Answers Ch 13
 - Setting Reading Goals Prentice Hall Chemistry Workbook Answers Ch 13
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Workbook Answers Ch 13
 - Fact-Checking eBook Content of Prentice Hall Chemistry Workbook Answers Ch 13
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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