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Name _____ Date _____ Class _____

17 THERMOCHEMISTRY

SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK (pages 505–510)

This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat.

► Energy Transformations (page 505)

1. What area of study in chemistry is concerned with the heat transfers that occur during chemical reactions? thermochemistry
2. Where the use of energy is concerned (in a scientific sense), when is work done?
   Work is done when a force is used to move an object.
3. Circle the letter next to each sentence that is true about energy.
   (a.) Energy is the capacity for doing work or supplying heat.
   (b.) Energy is detected only because of its effects.
   c. Heat is energy that transfers from one object to another because they are at the same temperature.
   (d.) Gasoline contains a significant amount of chemical potential energy.
4. Circle the letter next to each sentence that is true about heat.
   (a.) One effect of adding heat to a substance is an increase in the temperature of that substance.
   b. Heat always flows from a cooler object to a warmer object.
   (c.) If two objects remain in contact, heat will flow from the warmer object to the cooler object until the temperature of both objects is the same.

► Exothermic and Endothermic Processes (pages 506–507)

5. What can be considered the "system" and what are the "surroundings" when studying a mixture of chemicals undergoing a reaction? Write your answers where indicated below.
   System: The mixture of chemicals itself is considered the system.
   Surroundings: Everything but the mixture of chemicals is the surroundings, but practically speaking, the immediate vicinity of the system.

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Prentice Hall Chemistry Work Answers Chapter 19

Yan Bai

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