

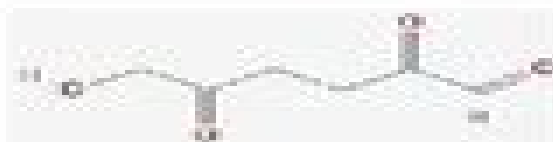
A) square planar B) bent see-saw C) trigonal pyramidal D) trigonal planar E)

Short Answer:



1a. How many pi bonds and sigma bonds are there in the molecule above? (2T)

# of sigma bonds in the molecule	
# of pi bonds in the molecule	



2b. What are the bond angles and hybridization of the oxygen of the O-H group in the molecule above? (2T)

question	angles	hybridization
2b.		

2c. What are the bond angles and hybridization of the carbon of the aldehyde in the molecule above? (2T)

question	angles	hybridization
2c.		

Prentice Hall Chemistry Covalent Bonding Test Answers

Terry C. Jones



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