

Chapter 7 Chemical Reactions

Section 7.1 Describing Reactions

(pages 192–198)

This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.

Reading Strategy (page 192)

Monitoring Your Understanding Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

Chemical Equations (pages 192–193)

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. true
- Circle the letter of each sentence that is a correct interpretation of the chemical equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$.
 - Carbon and oxygen react and form carbon monoxide.
 - ☒ Carbon and oxygen react and form carbon dioxide.
 - Carbon dioxide yields carbon and oxygen.
 - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. true
- Circle the letter of the correct answer. According to the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
 - 1
 - 7
 - ☒ 14
 - 28
- In the reaction represented by the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

Prentice Hall Chemistry 7 2 Answer

Ari L. Horvath



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Environmental Chemistry Stanley Manahan, Stanley E. Manahan, 2009-12-17 The field of environmental chemistry has evolved significantly since the publication of the first edition of Environmental Chemistry Throughout the book s long life it has chronicled emerging issues such as organochloride pesticides detergent phosphates stratospheric ozone depletion the banning of chlorofluorocarbons and greenhouse warming D *Gibbs Energy and Helmholtz Energy* Emmerich Wilhelm, Trevor Letcher, 2021-09-08 This book contains the latest information on all aspects of the most important chemical thermodynamic properties of Gibbs energy and Helmholtz energy as related to fluids Both the Gibbs energy and Helmholtz energy are very important in the fields of thermodynamics and material properties as many other properties are obtained from the temperature or pressure dependence Bringing all the information into one authoritative survey the book is written by acknowledged world experts in their respective fields Each of the chapters will cover theory experimental methods and techniques and results for all types of liquids and vapours This book is the fourth in the series of Thermodynamic Properties related to liquids solutions and vapours edited by Emmerich Wilhelm and Trevor Letcher The previous books were Heat Capacities 2010 Volume Properties 2015 and Enthalpy 2017 This book fills the gap in fundamental thermodynamic properties

and is the last in the series **Research and Development Progress Report** United States. Office of Saline Water,1965

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