

## Periodic Trends Worksheet

Directions: Use the table to determine the following properties.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

**K, Na, O, S, Cl**

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

**K, Na, S, O, Cl**

3. Which atom has the highest first ionization energy than sodium?

**Chlorine (because it has the most electronegativity and is closest to sodium)**

4. Which two elements in the same family generally have similar properties?

**There have been two atoms that share a same position with potassium**

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms of the greatest size? Of the smallest size listed?

**Period 7, period 1, increasing atomic number, decreasing size**

7. Which trend in ionization energy atoms are associated with the periodic table? Which trend is the opposite?

**Electron configuration, increasing electronegativity, decreasing atomic number**

8. Write the atoms in each pair that has the largest atomic radius.

a.	Br	Bi
b.	Na	As
c.	Al	Cl
d.	Fe	Si
e.	Al	Cl
f.	Na	As

a.	Bi	Br
b.	As	Na
c.	Cl	Al
d.	Si	Fe
e.	Cl	Al
f.	As	Na

9. Circle the atom in each pair that has the greatest electronegativity.

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**AW Chickering**



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## **Table of Contents Physical Science If8767 Pg 63 Key**

1. Understanding the eBook Physical Science If8767 Pg 63 Key
  - The Rise of Digital Reading Physical Science If8767 Pg 63 Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Pg 63 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science If8767 Pg 63 Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Pg 63 Key
  - Personalized Recommendations
  - Physical Science If8767 Pg 63 Key User Reviews and Ratings
  - Physical Science If8767 Pg 63 Key and Bestseller Lists
5. Accessing Physical Science If8767 Pg 63 Key Free and Paid eBooks

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  - Physical Science If8767 Pg 63 Key eBook Subscription Services
  - Physical Science If8767 Pg 63 Key Budget-Friendly Options
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- ePub, PDF, MOBI, and More
  - Physical Science If8767 Pg 63 Key Compatibility with Devices
  - Physical Science If8767 Pg 63 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science If8767 Pg 63 Key
  - Highlighting and Note-Taking Physical Science If8767 Pg 63 Key
  - Interactive Elements Physical Science If8767 Pg 63 Key
8. Staying Engaged with Physical Science If8767 Pg 63 Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science If8767 Pg 63 Key
9. Balancing eBooks and Physical Books Physical Science If8767 Pg 63 Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science If8767 Pg 63 Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science If8767 Pg 63 Key
- Setting Reading Goals Physical Science If8767 Pg 63 Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science If8767 Pg 63 Key
- Fact-Checking eBook Content of Physical Science If8767 Pg 63 Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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