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Date:

Class: Physikalische Chemie

Physikalische Chemie: Chapter 10: Outer Thermodynamic Energy and Work

Multiple Choice

1. The electrochemical potential
a. is defined as zero.
b. is arbitrary.
c. is always positive.
d. is always negative.
2. Energy added to a system
a. results mainly from external interactions.
b. is lost from the environment.
c. is transformed into heat.
d. increases the kinetic energy of the system.
3. When a definite quantity of work is done on an object,
a. mechanical work is done.
b. energy is lost from the system.
c. the object is 100% efficient.
d. it gains energy from whatever does the work.
4. A slow cyclic process (quasistatic process) usually is
a. irreversible.
b. reversible.
c. isothermal.
d. isobaric.
5. Spontaneous processes are
a. reversible.
b. irreversible.
c. always kinetic energy.
d. irreversible.
6. Why does liquid water freeze spontaneously?
a. condensation.
b. crystallization.
c. transformation.
d. evaporation.
7. Which statement is NOT true about heat?
a. It is a scalar quantity.
b. Heat always flows from high to low temperature.
c. Heat always flows from low to high temperature.
d. It is not a state function.
8. Which of the following is a thermal coordinate?
a. temperature.
b. heat.
c. work.
d. mass.

Physical Science Chapter 16 Section 1 Answers

Robert R. Redfield



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