

# Bonding and Reactions

## Study Guide

### 10 pages including KEY

### Chemical Bonding and Reactions Study Guide

This study guide will help you study for the Chemical Bonding and Reactions exam by organizing the most important concepts on structure, bonding and reactions. It is organized so that each section in the study guide corresponds to each section of the textbook and contains

#### B. VOCABULARY

Atom	Electron	Electron transfer	Electron transfer
Atomic mass	Electronegativity	Electronegativity	Electronegativity
Atomic number	Electron shell	Electron shell	Electronegativity
Chemical bond	Electron configuration	Electron shell	Electronegativity
Chemical structure	Electron configuration	Electron shell	Electronegativity

#### C. REASONING

1. Why do atoms bond together?
2. What type of bonds form?
3. What kind of elements form ionic bonds?
4. What is a covalent bond?
5. What kind of elements form covalent bonds?
6. Why don't noble gases form compounds with other elements?

Write the formula for the following compounds with multiple covalent bonds:

1. Oxygen difluoride \_\_\_\_\_

2. Oxygen difluoride \_\_\_\_\_

Write the following compounds with multiple covalent bonds:

3. Water \_\_\_\_\_

4. Methane \_\_\_\_\_

#### REACTIONS

How do you know when a reaction has occurred?

What does a chemical equation tell you about a reaction?

What is a coefficient in a reaction?

What is a subscript in a reaction?

How do you know the law of conservation of mass is obeyed?

### Bonding and Chemical Reactions - KEY

#### C. REASONING

1. Why do atoms bond together?  
Atoms bond together to achieve a more stable electronic configuration, usually by filling or emptying their valence electron shells to equal the number of electrons in the shells of noble gases.
2. What type of bonds form?  
Ions form ionic bonds. Atoms of different elements share electrons and form covalent bonds. Atoms of the same element share electrons and form covalent bonds.
3. What kind of elements form ionic bonds?  
Metals combine with nonmetals and form ionic bonds. Metals combine with nonmetals and form ionic bonds.
4. What is a covalent bond?  
A covalent bond is a bond formed by the sharing of electrons between two atoms. A covalent bond is a bond formed by the sharing of electrons between two atoms.
5. What kind of elements form covalent bonds?  
Nonmetals combine with nonmetals and form covalent bonds. Nonmetals combine with nonmetals and form covalent bonds.
6. Why don't noble gases form compounds with other elements?  
Noble gases have a full valence shell and are therefore very stable. They do not react with other elements.
7. How do you know when a reaction has occurred?  
A reaction has occurred when there is a change in color, temperature, or the formation of a precipitate or gas.
8. What does a chemical equation tell you about a reaction?  
A chemical equation tells you about the reactants and products of a reaction. It also tells you the relative amounts of each substance involved in the reaction.
9. What is a coefficient in a reaction?  
A coefficient is a number placed in front of a chemical formula in a balanced chemical equation to indicate the relative number of molecules of that substance involved in the reaction.
10. What is a subscript in a reaction?  
A subscript is a number placed below a chemical formula in a balanced chemical equation to indicate the number of atoms of that element in a molecule of that substance.
11. How do you know the law of conservation of mass is obeyed?  
The law of conservation of mass is obeyed when the total mass of the reactants is equal to the total mass of the products.
12. How do you know the law of conservation of mass is obeyed?  
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#### B. VOCABULARY

- Atom: A tiny particle consisting of protons, neutrons, and electrons.
- Atomic mass: The mass of an atom, usually expressed in atomic mass units.
- Atomic number: The number of protons in the nucleus of an atom.
- Chemical bond: The force that holds atoms together in a compound.
- Chemical structure: The arrangement of atoms in a molecule of a compound, showing how they are connected.
- Electron: A tiny negatively charged particle that is present in all atoms.
- Electronegativity: A measure of an atom's ability to attract electrons towards itself.
- Electron shell: A region of space around the nucleus of an atom where electrons are likely to be found.
- Electron transfer: The movement of electrons from one atom to another.
- Electron configuration: The arrangement of electrons in the shells of an atom.
- Electronegativity: A measure of an atom's ability to attract electrons towards itself.
- Electron shell: A region of space around the nucleus of an atom where electrons are likely to be found.
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# Physical Science Bonding And Chemical Reactions Test

**Yicheng Fang**



**Physical Science Bonding And Chemical Reactions Test:**

This book delves into Physical Science Bonding And Chemical Reactions Test. Physical Science Bonding And Chemical Reactions Test is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Physical Science Bonding And Chemical Reactions Test, encompassing both the fundamentals and more intricate discussions.

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    - Chapter 2: Essential Elements of Physical Science Bonding And Chemical Reactions Test
    - Chapter 3: Physical Science Bonding And Chemical Reactions Test in Everyday Life
    - Chapter 4: Physical Science Bonding And Chemical Reactions Test in Specific Contexts
    - Chapter 5: Conclusion
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