

Average Atomic Mass

How can the masses of the particles within an atom be determined?

Objectives

After observing three gases, three ions, and naturally occurring three isotopes, the masses of these particles, ions, and atoms should be determined. Comparison of particles including those belonging to the same element should show three different values resulting from isotopes of particles and measurement of ions should show three different values. The relative mass of protons and neutrons will be made apparent by an electron-mass measurement. The mass of the electron will be compared with the masses of the particles.

Model 1 - A Strip of Magnesium Metal



1. We see the distribution showing the mass. Mg comes in Model 1.
 Three gases show three different isotopes.
2. We see the masses of the particles showing three different isotopes (shown in Model 1).
 24, 25, 26.
3. We see the masses of the particles in Model 1. Three different masses are shown.
 The mass of the particles of the gases are different. The masses are 24, 25, 26.
4. We see the masses of the particles showing in Model 1. Three different masses are shown.
 The masses of the particles are different. The masses are 24, 25, 26.

	^{24}Mg	^{25}Mg	^{26}Mg
Abundance (%)	75	25	0

5. We see the masses of the particles in the same manner as Model 1.
 24, 25, 26.
 6. We see the masses of the particles in the same manner as Model 1. The masses of the particles are different. The masses are 24, 25, 26.
- As a group of 10 students, we will be approximately 24, 25, 26, and 27.

Periodicity Of Elements Pogil Key

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