

Graphing Periodic Trends Lab

Name _____ Hr _____

PURPOSE: To determine how certain properties are periodic when the elements are arranged in periods or groups.

DISCUSSION: There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

Ionization energy - The energy required to remove an electron from an atom.

DATA

Properties of the First 18 Elements

| Element | Atomic Number | Atomic Radius (Angstroms) | Ionization Energy |
|---------|---------------|---------------------------|-------------------|
| H | 1 | 0.79 | 314 |
| He | 2 | 0.49 | 567 |
| Li | 3 | 2.05 | 124 |
| Be | 4 | 1.4 | 215 |
| B | 5 | 1.17 | 191 |
| C | 6 | 0.91 | 260 |
| N | 7 | 0.75 | 335 |
| O | 8 | 0.65 | 314 |
| F | 9 | 0.57 | 402 |
| Ne | 10 | ? | ? |
| Na | 11 | 2.23 | 119 |
| Mg | 12 | 1.72 | 176 |
| Al | 13 | 1.82 | 138 |
| Si | 14 | 1.46 | 168 |
| P | 15 | 1.23 | 242 |
| S | 16 | 1.09 | 239 |
| Cl | 17 | 0.97 | 299 |
| Ar | 18 | ? | ? |

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with question marks.

1. Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)
2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)

Periodic Trends Graphing Lab Answers

R Sanford

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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