Periodic Trends Worksheet

Name: _____ ID: ____

Directions: Use your notes to answer the following questions.

Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Ans: 0, C, Al, K

Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Ans: Ne, Al, S, O

(3) Why does fluorine have a higher ionization energy than iodine?

Ans: Fluorine has a smaller atomic radius, so the proton can exert a greater pull on electron (Valence shell is close)

Why do elements in the same family generally have similar properties?

Ans: They have the same valence shell electron arrangement.

- 5 Indicate whether the following properties increase or decrease from left to right across the periodic table.
 - atomic radius (excluding noble gases) Decrease
 - b. first ionization energy Increase
 - electronegativity Increase
- (6) What trend in atomic radius occurs across the periodic table? What causes this trend?

Ans: Atomic radius increases down a group because energy levels (shells) are added.

What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ans: Ionization energy increases across a period because as elements become less metallic, it requires more energy to remove an electron.

8 Circle the atom in each pair that has the largest radius.

a. (Al) or B d. (O) or F

b. (Na) or Al e. (Br) or Cl

c. S or O f. Mg or Ca

Periodic Table Trends Answer Key

Pam Walker, Elaine Wood

Periodic Table Trends Answer Key:

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Table of Contents Periodic Table Trends Answer Key

- 1. Understanding the eBook Periodic Table Trends Answer Key
 - The Rise of Digital Reading Periodic Table Trends Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Table Trends Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Trends Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Table Trends Answer Key
 - Personalized Recommendations

- Periodic Table Trends Answer Key User Reviews and Ratings
- Periodic Table Trends Answer Key and Bestseller Lists
- 5. Accessing Periodic Table Trends Answer Key Free and Paid eBooks
 - o Periodic Table Trends Answer Key Public Domain eBooks
 - Periodic Table Trends Answer Key eBook Subscription Services
 - Periodic Table Trends Answer Key Budget-Friendly Options
- 6. Navigating Periodic Table Trends Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Periodic Table Trends Answer Key Compatibility with Devices
 - Periodic Table Trends Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Trends Answer Key
 - Highlighting and Note-Taking Periodic Table Trends Answer Key
 - Interactive Elements Periodic Table Trends Answer Key
- 8. Staying Engaged with Periodic Table Trends Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Trends Answer Key
- 9. Balancing eBooks and Physical Books Periodic Table Trends Answer Key
 - \circ Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Periodic Table Trends Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Trends Answer Key
 - Setting Reading Goals Periodic Table Trends Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Trends Answer Key
 - Fact-Checking eBook Content of Periodic Table Trends Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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