

Gas
Liquid
Solid
Unknown

Periodic Table

Cheat Sheet

Reactions

Single Displacement $\rightarrow A + BC = B + AC$
 Double Displacement $\rightarrow AB + CD = BC + AD$
 Synthesis $\rightarrow A + B = AB$
 Decomposition $\rightarrow AB = A + B$
 Combustion $\rightarrow \text{Hydrocarbon} + O_2 = CO_2 + H_2O$
 Neutralization $\rightarrow \text{Acid} + \text{Base} = \text{Salt} + H_2O$

Polyatomic ions (ion composed of multiple atoms)

$NO_3^- \rightarrow$ Nitrate (-1)

$PO_4^{3-} \rightarrow$ Phosphate (-3)

$OH^- \rightarrow$ Hydroxide (-1)

$HCO_3^- \rightarrow$ Bicarbonate (-1)

$NO_2^- \rightarrow$ Nitrite (-1)

$SO_4^{2-} \rightarrow$ Sulfate (-2)

$CO_3^{2-} \rightarrow$ Carbonate (-2)

$NH_4^+ \rightarrow$ Ammonium (+1)

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| Solid | | Periodic Table | | | | | | | | | | Liquid | | | | | | Gas | | | | | |
| Unknown | | Cheat Sheet | | | | | | | | | | Decomposition $\rightarrow A + B = A + B$ | | | | | | Combustion $\rightarrow \text{Hydrocarbon} + O_2 = CO_2 + H_2O$ | | | | | |
| | | | | | | | | | | | | Neutralization $\rightarrow \text{Acid} + \text{Base} = \text{Salt} + H_2O$ | | | | | | | | | | | |
| Polyatomic ions (ion composed of multiple atoms) | | | | | | | | | | | | | | | | | | | | | | | |
| $NO_3 \rightarrow \text{Nitrate } (-1)$ | | | | | | | | | | | | | | | | | | | | | | | |
| $PO_4 \rightarrow \text{Phosphate } (-3)$ | | | | | | | | | | | | | | | | | | | | | | | |
| $OH \rightarrow \text{Hydroxide } (-1)$ | | | | | | | | | | | | | | | | | | | | | | | |
| $HCO_3 \rightarrow \text{Bicarbonate } (-1)$ | | | | | | | | | | | | | | | | | | | | | | | |
| $NO_2 \rightarrow \text{Nitrite } (-1)$ | | | | | | | | | | | | | | | | | | | | | | | |
| $SO_4 \rightarrow \text{Sulfate } (-2)$ | | | | | | | | | | | | | | | | | | | | | | | |
| $CO_3 \rightarrow \text{Carbonate } (-2)$ | | | | | | | | | | | | | | | | | | | | | | | |
| $NH_4 \rightarrow \text{Ammonium } (+1)$ | | | | | | | | | | | | | | | | | | | | | | | |
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Alkali Metals:
 One valence electron
 Very reactive because they are ready to lose an electron to form cations
 Usually in ionic bonds due to reactivity
 Bonds with halogens, and other non-metals
 Have a vigorous reaction when bonding with halogens, oxygen and the compound of water

Alkaline Earth Metals:
 Two valence electrons
 Inactive but not as reactive as halogens and alkali metals because it needs to lose 2 electrons to have a full valence shell
 Bonds with halogens, and other non-metals

Transition Metals/ Metals: Transition metals are metals that vary in valence electrons and can have multiple ion charges. They have similar qualities as alkali metals and alkaline earth metals; they can form metallic and ionic bonds. Metals with multiple ion charges are called multivalent metals. For example, copper has an ion charge of +1 and +2 so it is a multivalent metal. Therefore, the ion charge must be specified when written in word form for ionic bonds with multivalent metals. Metals can also bond with polyatomic ions.

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|-----------------|---------------|--------------------|-----------------|------------------|-----------------|-----------------|------------------|-----------------|-------------------|-------------------|---------------|-------------------|-----------------|------------------|
| La Lanthanum | Ce Cerium | Pr Praseodymium | Nd Neodymium | Pm Promethium | Sm Samarium | Eu Europium | Gd Gadolinium | Tb Terbium | Dy Dysprosium | Ho Holmium | Er Erbium | Tm Thulium | Yb Ytterbium | Lu Lutetium |
| Ac Actinium | Th Thorium | Pa Protactinium | U Uranium | Np Neptunium | Pu Plutonium | Am Americium | Cm Curium | Bk Berkelium | Cf Californium | Es Einsteinium | Fm Fermium | Md Mendelevium | No Nobelium | Lr Lawrencium |

Metalloids: Contain metal and non-metal properties. They react with metals and non-metals to form covalent and ionic bonds but are less reactive than non-metals, depending on the electronegativity of the atoms.

Non-metals: They can not conduct electricity because they can not give off electrons, so current will not be able to go through.
 Non-metals bond with metals in an ionic bond and other non-metals in a molecular bond. They can also bond with the polyatomic ion ammonium due to its +1 charge.

Noble Gases: Non-metals in a previous form that have a full valence shell. This makes it not reactive, so it can not bond with other non-metals and metals. This is why noble gases are never in compounds and chemical reactions.

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Clemens Wendtner



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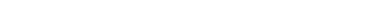
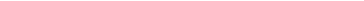
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