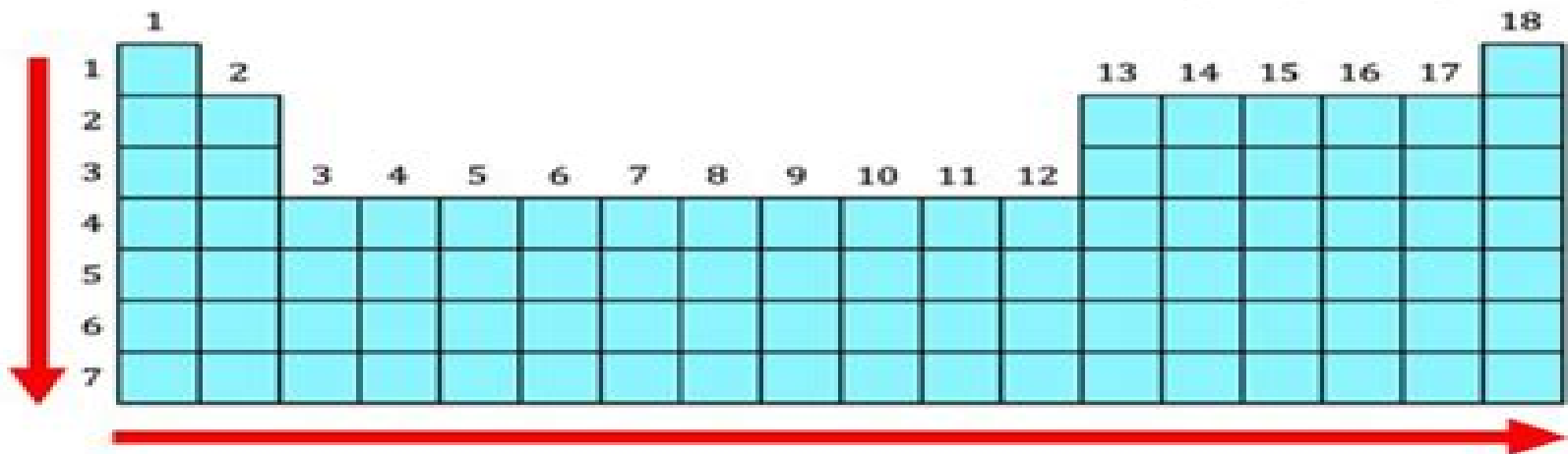


# Periodic Trends - Atomic Mass

- The **atomic mass** describes the mass of each individual element.
- **Atomic mass** increases moving **down** each group and **right** across a period.

**Instructions:** In the periodic table below, draw two arrows representing the direction that atomic mass **increases** across groups and periods.



**Instructions:** In each question, identify the element with the **greatest** atomic mass:

- 1) Chlorine (Cl), Iodine (I), Bromine (Br), Fluorine (F)      Iodine
- 2) Oxygen (O), Carbon (C), Nitrogen (N), Boron (B)      Oxygen
- 3) Lithium (Li), Silicon (Si), Sulfur (Br), Xenon (Xe)      Xenon

**Instructions:** In each question, list the elements from **least** to **greatest** atomic mass:

- 4) Magnesium (Mg), Beryllium (Be), Barium (Ba)      Be, Mg, Ba
- 5) Gallium (Ga), Selenium (Se), Potassium (K)      K, Ga, Se

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**Robert C. Smoot, Smoot, Richard G.  
Smith, Jack Price**

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**Online Searching** Joann H. Lee, 1989

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