

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. ☐ a. The modern periodic table, elements are ordered by according to the atomic number.
☐ b. according to Mendeleev's original design.
☒ c. according to increasing atomic number.
☐ d. based on how they were discovered.
2. ☐ a. Mendeleev stated that certain elements, under chemical properties of elements appeared at regular intervals which when discovered were arranged in order of increasing atomic number.
☐ b. elements.
☒ c. atomic number.
☐ d. increasing.
3. ☐ a. The modern periodic law states that
☐ b. all new elements with the same properties are listed in the same group.
☒ c. the physical and chemical properties of elements are functions of atomic number.
☐ d. elements exhibit properties of both periods and groups.
☐ e. the chemical properties of elements can be grouped according to periodicity but physical properties cannot.
4. ☐ a. The discovery of the noble gases changed Mendeleev's periodic table by adding a new
☐ b. period.
☐ c. series.
☒ d. group.
☒ e. level.
5. ☐ a. The noble gases that property of the noble gases is that they are
☐ b. unstable.
☐ c. inert.
☒ d. large molecules.
6. ☐ a. Likewise, the first column in Group 1, has an atomic number of 1. The second column in the group has an atomic number of
☐ b. 2.
☒ c. 10.
☐ d. 18.
7. How groups in the periodic table are numbered? (18 and the atomic number) of 10.
☐ a. How many periods are there are?
☐ b. How many elements are in this group?
☒ c. How many elements are in this group?
☐ d. How many elements are in this group?

Periodic Law Chapter 5 Test Answers

James E. Brady, Fred Senese



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