

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant

Excess: All reactants except the limiting reagent will have some left over. Remember: Limiting reagent always comes up first when figuring out.

Remember: Limiting reagent will determine how much product is produced and

byproducts will be. The limiting reagent is also a factor.



The ratio required for the components for the reaction is:

- (1) 10.00 g of H_2 and 10.00 g of O_2 will produce 10.00 g of H_2O .
Remember: Limiting reagent will be H_2 because it's the first one.



- (2) $\frac{10.00 \text{ g of } \text{H}_2}{2.016 \text{ g}} = 4.96$ (the limiting reagent) $\frac{10.00 \text{ g of } \text{O}_2}{32.00 \text{ g}} = 0.3125$ (the excess reagent)

- (3) 10.00 g of H_2 and 10.00 g of O_2 will produce 10.00 g of H_2O . Remember: Limiting reagent will be H_2 because it's the first one.



- (4) 10.00 g of H_2 and 10.00 g of O_2 will produce 10.00 g of H_2O . Remember: Limiting reagent will be H_2 because it's the first one.



Percent Yield And Limiting Reactants Quiz

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