Question	Answer	Marks	Guidance
1	Ca: N: O = 30,35/40.1: 21.20/14.0: 48.45/16.0 = 0.7569: 1.5286: 3.028	81	
	Formula = CaN ₂ O ₄	B1	
2 (a)	2KClO ₃ (s) → 2KCl(s) + 3O ₂ (g)	B1	
2 (b)	4.50 × 10 ⁻³ mol	B1	
2 (c)	$n(KCIO_3) = 3.0 \times 10^{-3} \text{ mol}$	B1	
	$M(KCIO_3) = 122.6 \text{ g mol}^{-1}$	B1	
	mass of KCIO ₃ = 0.3678 g	B1	
3	M(ZnSO ₄) = 161.5 g mol ⁻¹	B1	
	$n(ZnSO_4) = 6.57 \times 10^{-3} \text{ mol}$	B1	
	$n(H_2O) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$ $x = n(H_2O)/n(ZnSO_4) = 7$	B1	
	Formula = ZnSO ₄ •7H ₂ O	B1	
4	$M(Na_2CO_3) = 106.0 \text{ g mol}^{-1}$	B1	
	Actual n(Na ₂ CO ₃) = 0.0234 mol	B1	
	Theoretical ri(Na ₂ CO ₃) = 0.0234 × 100/65 = 0.0360 mol	B1	
	Theoretical n(NaHCO ₃) = 0.0720 mol	B1	
	Mass of NaHCO ₃ = 0.0720 × 84.0 = 6.05 g	B1	
5 (a)	$Fe_2O_3(s) + 3CO(g) \rightarrow 2Fe(s) + 3CO_2(g)$	B1	
5 (b)	$M(Fe_2O_3) = 159.6 \text{ g mol}^{-1}$	B1	
	n(Fe ₂ O ₃) = (10 000 × 10 ⁶)159.6 = 6.266 × 10 ⁷ mol	B1	
	n(Fe) = 1.253 × 10 ⁸ mol = 6.992 × 10 ⁹ g (6992 tonne)	B1	
6 (a)	Na : N : O = 27.1/23.0 : 16.5/14.0 : 56.4/16.0 = 1.178 : 1.179 : 3.525	B1	
	Formula = NaNO ₃	B1	
6 (b)	2NaNO₂(s) → 2NaNO₂(s) + O₂(g)	B1	

Ocr Chemistry Exam Question Answers

Mike Smith

Ocr Chemistry Exam Question Answers:

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Table of Contents Ocr Chemistry Exam Question Answers

- 1. Understanding the eBook Ocr Chemistry Exam Question Answers
 - The Rise of Digital Reading Ocr Chemistry Exam Question Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Ocr Chemistry Exam Question Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry Exam Question Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ocr Chemistry Exam Question Answers
 - Personalized Recommendations
 - Ocr Chemistry Exam Question Answers User Reviews and Ratings
 - Ocr Chemistry Exam Question Answers and Bestseller Lists
- 5. Accessing Ocr Chemistry Exam Question Answers Free and Paid eBooks

- Ocr Chemistry Exam Question Answers Public Domain eBooks
- Ocr Chemistry Exam Question Answers eBook Subscription Services
- o Ocr Chemistry Exam Question Answers Budget-Friendly Options
- 6. Navigating Ocr Chemistry Exam Question Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - o Ocr Chemistry Exam Question Answers Compatibility with Devices
 - Ocr Chemistry Exam Question Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry Exam Question Answers
 - Highlighting and Note-Taking Ocr Chemistry Exam Question Answers
 - Interactive Elements Ocr Chemistry Exam Question Answers
- 8. Staying Engaged with Ocr Chemistry Exam Question Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry Exam Question Answers
- 9. Balancing eBooks and Physical Books Ocr Chemistry Exam Question Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry Exam Question Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Chemistry Exam Question Answers
 - Setting Reading Goals Ocr Chemistry Exam Question Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Chemistry Exam Question Answers
 - Fact-Checking eBook Content of Ocr Chemistry Exam Question Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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