## Chemical Nomenclature Review

#### Introduction

This is a worksheet reviewing some of the important concepts you will need to use from previous chapters to be successful with stoichiometry. Each section has a sample problem followed by a series of practice questions.

#### Ionic Nomenclature

**Ions** are atoms that have a charge because they have gained or lost electrons. **Ionic compounds** are formed when two ions of opposite charges are attracted to each other and form a bond. To name them, simply write the name of the cation (positive ion) followed by the name of the anion (negative ion).

Here is an example of a simple binary ionic compound. The anion in these compounds ends with -ide:

CaBr. Calcium bromide

Polyatomic ions are made of multiple atoms. The anion will usually end with -ate or -ite.

NaNO<sub>3</sub>

Sodium nitrate

Any compound with a transition meta	al needs a Roman numeral to show the charge of the cation.
Zr <sub>2</sub> O <sub>3</sub>	Zirconium (III) oxide

=1.1	MgS	Magnesium Sulfide	942	NH <sub>4</sub> CI	Ammonium chloride
2.	KF	Potassium fluoride	5.	TiO <sub>2</sub>	Titanium dioxide
3.	$Ca_3(PO_4)_2$	Calcium phosphate	6.	$Cu_2CO_3$	carbon (II) carbonate hydroxide

#### Ionic Formulas

Ionic formulas must be written in such a way that the charges of the cation and anion balance and form a neutral compound. This can be done by adding **subscripts** to change the number of each ion within the compound.

45	- 1-20 Contraction	
Cesiun	n nitride	Cs₃N

Remember, Roman numerals after transition metals indicate the charge of the cation.

		Zinc (II) hydroxide	Zn(OH) <sub>2</sub>
7.	Calcium chloride	Ca/Cl <sub>1</sub>	10. Iron (III) chlorite FeCla
8.	Aluminum phosphide	AIP	11. Silver (I) oxalate Ag:C:O:
9.	Potassium carbonate	K:CO:	12. Zinc (II) borate B <sub>2</sub> O <sub>4</sub> Zn <sub>1</sub>

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