

b) How many grams of phosphorus are required to react completely with 20.0 g of  $O_2$ ?

$$20.0 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{31.998 \text{ g } O_2} \times \frac{4 \text{ mol P}}{5 \text{ mol } O_2} \times \frac{30.973 \text{ g P}}{1 \text{ mol P}} = 15.5 \text{ g P}$$

c) What is the theoretical yield in grams if you combine the amounts of reactants in part b?

$$20.0 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{31.998 \text{ g } O_2} \times \frac{2 \text{ mol } P_2O_5}{5 \text{ mol } O_2} \times \frac{141.941 \text{ g } P_2O_5}{1 \text{ mol } P_2O_5} = 35.5 \text{ g } P_2O_5$$

OR

$$15.5 \text{ g P} \times \frac{1 \text{ mol P}}{30.973 \text{ g P}} \times \frac{2 \text{ mol } P_2O_5}{4 \text{ mol P}} \times \frac{141.941 \text{ g } P_2O_5}{1 \text{ mol } P_2O_5} = 35.5 \text{ g } P_2O_5$$

5. A gaseous mixture containing 5.00 moles of  $H_2$  and 7.00 moles of  $Br_2$  reacts to form  $HBr$ .

a) Write a balanced chemical equation for this reaction.



b) Which reactant is limiting?

1 to 1 ratio  
5.00 mol  $H_2$  & 7.00 mol  $Br_2$   
 $\therefore H_2$  is limiting

c) What is the theoretical yield for this reaction in moles?

$$5.00 \text{ mol } H_2 \times \frac{2 \text{ mol } HBr}{1 \text{ mol } H_2} = 10.0 \text{ mol } HBr$$

d) What is the theoretical yield for this reaction in grams?

$$10.0 \text{ mol } HBr \times \frac{80.912 \text{ g } HBr}{1 \text{ mol } HBr} = 809 \text{ g } HBr$$

6. Sodium reacts with molecular hydrogen to form sodium hydride. A reaction mixture contains 10.0 g Na and 0.0235 g  $H_2$ .

a) Write a balanced chemical equation for this reaction.



b) Which reactant is limiting?  $H_2$  is limiting

$$10.00 \text{ g Na} \times \frac{1 \text{ mol Na}}{22.9898 \text{ g Na}} = 0.4350 \text{ mol Na} \quad 0.0235 \text{ g } H_2 \times \frac{1 \text{ mol } H_2}{2.0158 \text{ g } H_2} = 0.0117 \text{ mol } H_2$$
$$\frac{0.4350}{2} = 0.2175 \quad \frac{0.0117}{1} = 0.0117$$

c) What is the theoretical yield for this reaction in grams?

$$0.0117 \text{ mol } H_2 \times \frac{2 \text{ mol NaH}}{1 \text{ mol } H_2} \times \frac{23.998 \text{ g NaH}}{1 \text{ mol NaH}} = 0.562 \text{ g NaH}$$

d) When this reaction is actually performed, 0.428 g of  $NaH$  is recovered. What is the percent yield of the reaction?

$$\frac{0.428 \text{ g}}{0.562} \times 100 = 76.2\%$$

# Periodic Trends Pogil Extension Questions

**R Pring**



## **Periodic Trends Pogil Extension Questions:**

## **Periodic Trends Pogil Extension Questions** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Periodic Trends Pogil Extension Questions**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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