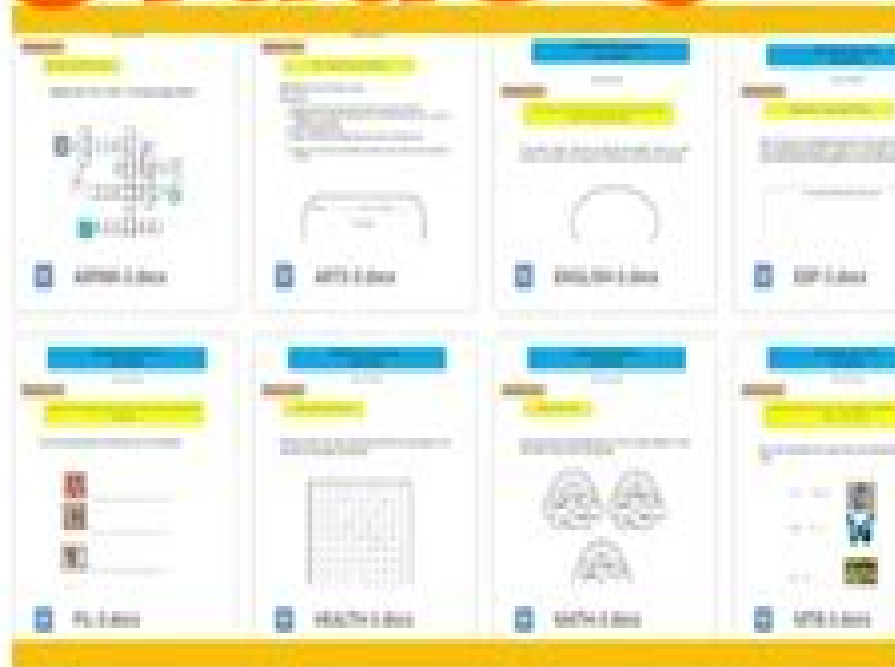


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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

april 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} \rightarrow 2\text{Cl}^- \text{aq}$

ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{ Ag}^+ \text{aq} + 2 \text{ F}^- \text{aq} + 2 \text{ NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2 \text{ NH}_4^+ \text{aq} + 2 \text{ F}^- \text{aq}$ because two $\text{NH}_4^+ \text{aq}$

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