

- (c) Using the redox systems in **Table 6.1**, construct an equation for a reaction between acidified dichromate(VI) ions and methanoic acid, HCOOH .

Rather than using [O] or [H], your equation must show the actual reactants and products.

[2]

- (d) A student added some chromium metal to an acidified solution containing copper(II) ions. A reaction took place. The student concluded that 'chromium is more reactive than copper'.

- (i) Explain, in terms of their electrode potentials, why 'chromium is more reactive than copper' in this reaction.

.....

 [2]

- (ii) When this experiment was carried out, the student observed some bubbles of a gas. Suggest an explanation for this observation.

.....
 [1]

- (e) Methanoic acid, HCOOH , has the common name of 'formic acid'. Direct-Formic Acid Fuel Cells (DFAFCs) are being developed for use in small, portable electronics such as phones and laptop computers.

In this fuel cell, methanoic acid (the fuel) reacts with oxygen to generate a cell potential.

- (i) Predict the standard cell potential of a DFAFC.

standard cell potential = V [1]

- (ii) Suggest **two** advantages of using methanoic acid as the fuel in a fuel cell rather than hydrogen.

.....

 [2]

[Total: 14]

Turn over

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