

(d) In an experiment, a student makes a solution of strontium chloride, SrCl_2 , by adding excess dilute hydrochloric acid to strontium carbonate.

(i) Describe what the student would observe and write the equation for the reaction.

observations

.....

equation [2]

(ii) Draw a 'dot-and-cross' diagram to show the bonding of strontium chloride. Show **outer** electrons only.

[2]

(e) In another experiment, a student attempts to make a solution of strontium chloride by adding chlorine water to aqueous strontium bromide.

(i) Describe what the student would observe.

..... [1]

(ii) Write the ionic equation for the reaction which takes place.

..... [1]

(iii) Chlorine is more reactive than bromine. Explain why.

.....

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.....

..... [4]

[Total: 18]

Turn over

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Table of Contents Ocr Decision 1 Jan 13 Paper

1. Understanding the eBook Ocr Decision 1 Jan 13 Paper
 - The Rise of Digital Reading Ocr Decision 1 Jan 13 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Decision 1 Jan 13 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Decision 1 Jan 13 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Decision 1 Jan 13 Paper
 - Personalized Recommendations

- Ocr Decision 1 Jan 13 Paper User Reviews and Ratings
- Ocr Decision 1 Jan 13 Paper and Bestseller Lists
- 5. Accessing Ocr Decision 1 Jan 13 Paper Free and Paid eBooks
 - Ocr Decision 1 Jan 13 Paper Public Domain eBooks
 - Ocr Decision 1 Jan 13 Paper eBook Subscription Services
 - Ocr Decision 1 Jan 13 Paper Budget-Friendly Options
- 6. Navigating Ocr Decision 1 Jan 13 Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Decision 1 Jan 13 Paper Compatibility with Devices
 - Ocr Decision 1 Jan 13 Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Decision 1 Jan 13 Paper
 - Highlighting and Note-Taking Ocr Decision 1 Jan 13 Paper
 - Interactive Elements Ocr Decision 1 Jan 13 Paper
- 8. Staying Engaged with Ocr Decision 1 Jan 13 Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Decision 1 Jan 13 Paper
- 9. Balancing eBooks and Physical Books Ocr Decision 1 Jan 13 Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Decision 1 Jan 13 Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Decision 1 Jan 13 Paper
 - Setting Reading Goals Ocr Decision 1 Jan 13 Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Decision 1 Jan 13 Paper
 - Fact-Checking eBook Content of Ocr Decision 1 Jan 13 Paper

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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