

16. The dissociation of water is shown below:



- At 25 °C, the auto product of water, K_w , is 1.0×10^{-14} mol² dm⁻⁶.
At 65 °C, the auto product of water, K_w , is 1.0×10^{-13} mol² dm⁻⁶.

17. Explain whether the dissociation of water is an endothermic or exothermic process.
Endothermic

18. Predict, using a calculation, whether a solution of 1 dm³ of 0.1 M is neutral, acidic or alkaline. (3)

Acidic

19 Explain by calculating pH.

20. K_w , pK_w and pH are significant notions.

Calculate pK_w at 65 °C.

Give your answer to two decimal places.

$$-\log(1.0 \times 10^{-13}) = 13.00$$

$$pK_w = 13.00 \quad (1)$$

21. Describe, at least two ways, how to obtain pure water from sea water without using any chemical reagent at all.

Calculate the pH of the resulting solution of water at 65 °C.

Give your answer to two decimal places.

$$pH = 7.00$$

pH = _____ (1)
(Total: 10)

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