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4.14.

$$\left(\frac{\partial H_m}{\partial T}\right)_P = \frac{C_{P,m}}{T} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\text{and } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) \left(\frac{1 \text{ mol}}{1000 \text{ g}} \right) = 1.25 \text{ J K}^{-1} \text{ g}^{-1}$$

and that we can estimate $C_{V,m}$ as follows:

- 4.15. (a) The weak van der Waals force between the atoms is positive in sign.
 (b) The weak van der Waals force between the atoms is positive in sign.
 (c) The weak van der Waals force between the atoms is positive in sign.

4.16. From the data we find the heat capacity of the gas is $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$ and that the heat capacity of the solid is

$$C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) \left(\frac{1 \text{ mol}}{1000 \text{ g}} \right) = 1.25 \text{ J K}^{-1} \text{ g}^{-1}$$

4.17. From the data we find $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$ and that the heat capacity of the solid is

$$C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) \left(\frac{1 \text{ mol}}{1000 \text{ g}} \right) = 1.25 \text{ J K}^{-1} \text{ g}^{-1}$$

$$C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) \left(\frac{1 \text{ mol}}{1000 \text{ g}} \right) = 1.25 \text{ J K}^{-1} \text{ g}^{-1}$$

4.18. From the data we find the heat capacity of the solid is

$$C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) \left(\frac{1 \text{ mol}}{1000 \text{ g}} \right) = 1.25 \text{ J K}^{-1} \text{ g}^{-1}$$

Physical Chemistry Ball Solution Manual

G Psacharopoulos



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