| | | gerangskik sis om die hoeke tussen hulle te maksimeer. Strukturele elektronpare is bindingspare plus alleenpare. OF Valensdop-elektronpaarafstotingsmodel is 'n model om die vorms van molekules te voorspels waarin die elektronpaar vanaf die buitenste dops van 'n verwysingsatoom rondom hierdie atoom gerangskik is om die afstoting tussen hulle te minimaliseer. | |
|-----|-------|--|-----|
| | 1.2.4 | The molecule is a tetrahedron [™] Die molekule is 'n tetraëder [™] 1 mark/punt | LO2 |
| | 1.2.5 | Name trigonal pyramid Noem trigonale piramidale 3 marks/punte | LO4 |
| 1.3 | 1.3.1 | The electron pair is equally shared between the two CI atoms and the resulting Cl ₂ molecule has no degree of polarity. Die elektronpaar word ewe veel gedeel tussen die twee CI-atome en die resulterende CI2-molekule het geen mate van polariteit nie. 2 marks/punte | LO2 |
| | 1.3.2 | The tendency of an atom to pull the bonded pair of electrons towards itself. of Die neiging of van 'n atoom om die gebinde elektronpare na homself te trek of a marks/punte | LO2 |
| | 1.3.3 | Electronegativity of Na = 6 Electronegativity of Cl = 3 Difference is 3 With a difference of greater than 2.1 electron transfer will take place and the bond elis ionic. (value of 2.1 does not need to be given for this mark) Elektronegatiwiteit van Na = 6 Elektronegatiwiteit van Cl = 3 Verskil is 3 Met 'n verskil van meer as 2,1 sal elektronoordrag plaasvind en die binding is ionies. (waarde van 2.1 hoef nie vir hierdie punt gegee te word nie) 4 marks/punte | LO2 |

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Karin Nielsen-Saines

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Table of Contents Physical Science Memo2013 Control Grade 1september

- 1. Understanding the eBook Physical Science Memo2013 Control Grade 1september
 - The Rise of Digital Reading Physical Science Memo2013 Control Grade 1september
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Memo2013 Control Grade 1september
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Memo2013 Control Grade 1september
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Memo2013 Control Grade 1september
 - Personalized Recommendations

- Physical Science Memo2013 Control Grade 1september User Reviews and Ratings
- Physical Science Memo2013 Control Grade 1september and Bestseller Lists
- 5. Accessing Physical Science Memo2013 Control Grade 1september Free and Paid eBooks
 - Physical Science Memo2013 Control Grade 1september Public Domain eBooks
 - Physical Science Memo2013 Control Grade 1september eBook Subscription Services
 - Physical Science Memo2013 Control Grade 1september Budget-Friendly Options
- 6. Navigating Physical Science Memo2013 Control Grade 1september eBook Formats
 - o ePub, PDF, MOBI, and More
 - Physical Science Memo2013 Control Grade 1september Compatibility with Devices
 - Physical Science Memo2013 Control Grade 1september Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Memo2013 Control Grade 1september
 - Highlighting and Note-Taking Physical Science Memo2013 Control Grade 1september
 - Interactive Elements Physical Science Memo2013 Control Grade 1september
- 8. Staying Engaged with Physical Science Memo2013 Control Grade 1september
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Memo2013 Control Grade 1september
- 9. Balancing eBooks and Physical Books Physical Science Memo2013 Control Grade 1september
 - ∘ Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Memo2013 Control Grade 1september
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Memo2013 Control Grade 1september
 - Setting Reading Goals Physical Science Memo2013 Control Grade 1september
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Memo2013 Control Grade 1september
 - Fact-Checking eBook Content of Physical Science Memo2013 Control Grade 1september

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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