

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Post Lab Answers Periodic Trends

Michael Brown



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Table of Contents Post Lab Answers Periodic Trends

1. Understanding the eBook Post Lab Answers Periodic Trends
 - The Rise of Digital Reading Post Lab Answers Periodic Trends
 - Advantages of eBooks Over Traditional Books
2. Identifying Post Lab Answers Periodic Trends
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Post Lab Answers Periodic Trends
 - User-Friendly Interface
4. Exploring eBook Recommendations from Post Lab Answers Periodic Trends
 - Personalized Recommendations
 - Post Lab Answers Periodic Trends User Reviews and Ratings
 - Post Lab Answers Periodic Trends and Bestseller Lists
5. Accessing Post Lab Answers Periodic Trends Free and Paid eBooks

- Post Lab Answers Periodic Trends Public Domain eBooks
 - Post Lab Answers Periodic Trends eBook Subscription Services
 - Post Lab Answers Periodic Trends Budget-Friendly Options
6. Navigating Post Lab Answers Periodic Trends eBook Formats
- ePub, PDF, MOBI, and More
 - Post Lab Answers Periodic Trends Compatibility with Devices
 - Post Lab Answers Periodic Trends Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Post Lab Answers Periodic Trends
 - Highlighting and Note-Taking Post Lab Answers Periodic Trends
 - Interactive Elements Post Lab Answers Periodic Trends
8. Staying Engaged with Post Lab Answers Periodic Trends
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Post Lab Answers Periodic Trends
9. Balancing eBooks and Physical Books Post Lab Answers Periodic Trends
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Post Lab Answers Periodic Trends
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Post Lab Answers Periodic Trends
- Setting Reading Goals Post Lab Answers Periodic Trends
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Post Lab Answers Periodic Trends
- Fact-Checking eBook Content of Post Lab Answers Periodic Trends
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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