

CHEM 101 Quiz #6 Chapter 6 ANSWERS**1. (4 points)**

The specific heat (s) of iron (Fe) is 0.40 J/(g K) . Calculate the heat effect q when 20 g of iron is heated from 100°C to 200°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \Delta K = \Delta^\circ\text{C} \quad q = 20 \times 0.40 \times (200 - 100) = +800 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Mg = 24.0 g/mol .

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

$$2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol}) \times (-600)(\text{kJ/mol}) = -60 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following four reactions is a *formation reaction*? Circle the correct answer.



Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur so e.g. Cl_2 not Cl , and H_2 not H !

CHEM 101 Section L52 Quiz #6b**1. (4 points)**

The specific heat (s) of copper (Cu) is 0.38 J/(g K) . Calculate the heat effect q when 10 g of copper is heated from 100°C to 300°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \Delta K = \Delta^\circ\text{C} \quad q = 10 \times 0.38 \times (300 - 100) = +760 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Ca = 40.0 g/mol .

Calculate the enthalpy change ΔH when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

$$8.0 \text{ g Ca} = 8.0/40 = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol}) \times (-650)(\text{kJ/mol}) = -130 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following reactions is a *formation reaction*? Circle the correct answer.



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Table of Contents Review Sheet Unit Answers Chemistry Answers

1. Understanding the eBook Review Sheet Unit Answers Chemistry Answers
 - The Rise of Digital Reading Review Sheet Unit Answers Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Sheet Unit Answers Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Sheet Unit Answers Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Sheet Unit Answers Chemistry Answers
 - Personalized Recommendations
 - Review Sheet Unit Answers Chemistry Answers User Reviews and Ratings
 - Review Sheet Unit Answers Chemistry Answers and Bestseller Lists

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 - Highlighting and Note-Taking Review Sheet Unit Answers Chemistry Answers
 - Interactive Elements Review Sheet Unit Answers Chemistry Answers
8. Staying Engaged with Review Sheet Unit Answers Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Sheet Unit Answers Chemistry Answers
9. Balancing eBooks and Physical Books Review Sheet Unit Answers Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Sheet Unit Answers Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Review Sheet Unit Answers Chemistry Answers
 - Setting Reading Goals Review Sheet Unit Answers Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Review Sheet Unit Answers Chemistry Answers
 - Fact-Checking eBook Content of Review Sheet Unit Answers Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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