Name: \_\_\_\_\_\_ Date: \_\_\_\_\_\_

# **Net Ionic Equations: Practice Problems**

Balance the following equations. Show the ionic and net ionic forms.

- Mg(NO<sub>3</sub>)<sub>2</sub>(aq.) + Na<sub>2</sub>CrO<sub>4</sub>(aq.) → MgCrO<sub>4</sub>(s) + NaNO<sub>3</sub>(aq.) lonic:
   Net ionic:
- 2) Na<sub>3</sub>PO<sub>4</sub>(aq.) + Ni(ClO<sub>4</sub>)<sub>2</sub>(s) → Ni<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>(s) + NaClO<sub>4</sub>(aq.) lonic:
  Net ionic:
- 3) AgNO<sub>3</sub>(aq.) + MgI<sub>2</sub>(aq.) → AgI(s) + Mg(NO<sub>3</sub>)<sub>2</sub>(aq.) lonic:
  Net ionic:
- 4) Ca(OH)₂(aq.) + FeCl₃(aq.) → Fe(OH)₃(s) + CaCl₂(aq.) lonic:
  Net ionic:
- 5) RbF(aq.) + CuSO₄(aq.) → Rb₂SO₄(aq.) + CuF₂(aq.) lonic:
  Net ionic:
- 6) (NH<sub>4</sub>)<sub>2</sub>S(aq.) + Co(NO<sub>3</sub>)<sub>2</sub>(aq.) → CoS(s) + NH<sub>4</sub>NO<sub>3</sub>(aq.) lonic:
  Net ionic:
- 7) Na<sub>3</sub>PO<sub>4</sub>(aq.) + FeCl<sub>3</sub>(aq.) → FePO<sub>4</sub>(s) + NaCl(aq.) lonic:
  Net ionic:
- 8) AgNO<sub>3</sub>(aq.) + HCI(aq.) → AgCI(s) + HNO<sub>3</sub>(aq.) lonic:
  Net ionic:
- 9) MgBr₂(aq.) + KOH(aq.) → Mg(OH)₂(s) + KBr(aq.) lonic: Net ionic:
- 10) AlCl<sub>3</sub>(aq.) + K<sub>3</sub>PO<sub>4</sub>(aq.) → AlPO<sub>4</sub>(s) + KCl(aq.) lonic:
  Net ionic:

# **Questions On Net Ionic Equations**

Neil D. Jespersen, Alison Hyslop

**Questions On Net Ionic Equations:** 

The book delves into Questions On Net Ionic Equations. Questions On Net Ionic Equations is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and indepth insights into Questions On Net Ionic Equations, encompassing both the fundamentals and more intricate discussions.

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