

Chapter 4 - Section A - Mathcad Solutions

- 4.1 (a) $T_0 := 473.15 \cdot \text{K}$ $T := 1373.15 \cdot \text{K}$ $n := 10 \cdot \text{mol}$
 For SO_2 : $A := 5.699$ $B := 0.801 \cdot 10^{-3}$ $C := 0.0$ $D := -1.015 \cdot 10^{-5}$
 $\Delta H := R \cdot \text{ICPH}(T_0, T, A, B, C, D)$
 $\Delta H = 47.007 \frac{\text{kJ}}{\text{mol}}$ $Q := n \cdot \Delta H$
 $Q = 470.073 \text{ kJ}$ Ans.
- (b) $T_0 := 523.15 \cdot \text{K}$ $T := 1473.15 \cdot \text{K}$ $n := 12 \cdot \text{mol}$
 For propane $A := 1.213$ $B := 28.785 \cdot 10^{-3}$ $C := -8.824 \cdot 10^{-6}$ $D := 0$
 $\Delta H := R \cdot \text{ICPH}(T_0, T, A, B, C, 0.0)$
 $\Delta H = 161.834 \frac{\text{kJ}}{\text{mol}}$ $Q := n \cdot \Delta H$
 $Q = 1.942 \times 10^3 \text{ kJ}$ Ans.
- 4.2 (a) $T_0 := 473.15 \cdot \text{K}$ $n := 10 \cdot \text{mol}$ $Q := 800 \cdot \text{kJ}$
 For ethylene: $A := 1.424$ $B := \frac{14.394 \cdot 10^{-3}}{\text{K}}$ $C := \frac{-4.392 \cdot 10^{-6}}{\text{K}^2}$
 $\tau := 2$ (guess) Given
 $Q = n \cdot R \cdot \left[\left[A \cdot T_0 \cdot (\tau - 1) + \frac{B}{2} \cdot T_0^2 \cdot (\tau^2 - 1) \right] + \frac{C}{3} \cdot T_0^3 \cdot (\tau^3 - 1) \right]$
 $\tau := \text{Find}(\tau)$ $\tau = 2.905$ $T := \tau \cdot T_0$ $T = 1374.5 \text{ K}$ Ans.
- (b) $T_0 := 533.15 \cdot \text{K}$ $n := 15 \cdot \text{mol}$ $Q := 2500 \cdot \text{kJ}$
 For 1-butene: $A := 1.967$ $B := \frac{31.630 \cdot 10^{-3}}{\text{K}}$ $C := \frac{-9.873 \cdot 10^{-6}}{\text{K}^2}$

Quation And Soultion Manual For Nuclear Physics

L Towne



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