

n l m_l m_s

1. State the four quantum numbers and the possible values they may have.

First - Primary Quantum number (n) = size of electron cloud n = 1 up to ∞....in reality n = 1 -7

Second – Azimuthal or Angular Momentum Quantum number (l) = shape of electron cloud;

l = 0 up to (n-1).... When l = 0 (s cloud), 1 (p cloud), 2 (d cloud), 3 (f cloud)

Third – Magnetic Quantum number (m_l) = location or spatial orientation m_l = - l to l

Fourth – Spin Quantum number (m_s) = electron spin m_s = +1/2 or -1/2

2. Name the orbitals described by the following quantum numbers

a. n = 3, l = 0

3s

c. n = 3, l = 2

3d

b. n = 3, l = 1

3p

d. n = 5, l = 0

5s

3. Give the n and l values for the following orbitals

a. 1s

(1,0)

d. 4d

(4,2)

b. 3s

(3,0)

e. 5f

(5,3)

c. 2p

(2,1)

4. Place the following orbitals in order of increasing energy:

1s, 3s, 4s, 6s, 3d, 4f, 3p, 7s, 5d, 5p **1s 3s 3p 4s 3d 5p 6s 4f 5d 7s**

5. What are the possible m_l values for the following types of orbitals?

a. s

0

c. d

-2, -1, 0, +1, +2

b. p

-1, 0, +1

d. f

-3, -2, -1, 0, +1, +2, +3

6. How many possible orbitals are there for n =

a. 4

16 = n² 4s 4p_x, 4p_y, 4p_z 4d_x, 4d_y, 4d_z, 4d_{xy} 4f_x, 4f_y, 4f_z, 4f_{xy}, 4f_{yz}, 4f_{xz}

b. 10

100

7. How many electrons can inhabit all of the n=4 orbitals?

32 = 2n² 4s² 4p_x², 4p_y², 4p_z² 4d_x², 4d_y², 4d_z², 4d_{xy}² 4f_x², 4f_y², 4f_z², 4f_{xy}², 4f_{yz}², 4f_{xz}²

Identify the element whose last electron would have the following four quantum numbers:

8. 3, 1, -1, +1/2 **Al**

11. 4, 3, +3, -1/2 **Lu**

Quantum Numbers Practice Questions Answers

Camilla Rothe



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