

Multiple Choice: Quantum Numbers

1. The Principal quantum number (n) of an electron indicates the

- A) energy level (shell) of the electron B) mass of the electron
C) charge of the electron D) spin of the electron

ANS: A

1. Which of the following is always true of all electrons that have the same principal quantum number.

- A) they possess the same energy B) they have opposite spins C) they are in identical orbitals
D) all of the above E) none of the above

ANS: E

1. An electron whose quantum numbers are $n = 3$, $l = 2$ and $m_l = 1$, is found in

- A) an s orbital B) a p orbital C) a d orbital D) an f orbital

ANS: C

1. Which of the following describes a $4p$ electron?

- A) $(4,1,0,1/2)$ B) $(4,0,0,1/2)$ C) $(3,2,2,1/2)$ D) $(3,2,-2,-1/2)$

ANS: A

1. If $l=1$, which of the following is not a possible value for m_l ?

- A) 0 B) 1 C) -1 D) 2

ANS: D

1. The _____ quantum number defines the shape of an orbital.

- A) spin B) magnetic C) principal D) secondary E) phi

ANS: D

1. The azimuthal quantum number is 3 in _____ orbitals.

- A) s B) p C) d D) f E) g

ANS: D

1. The principal quantum number of the first d subshell is _____.

- A) 1 B) 2 C) 3 D) 4 E) 5

ANS: C

Quantum Chemistry Multiple Choice Questions

R Sanford



Quantum Chemistry Multiple Choice Questions:

Decoding **Quantum Chemistry Multiple Choice Questions**: Revealing the Captivating Potential of Verbal Expression

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