

# CHEMISTRY I EQUATION SHEET

## Solubility Rules

| Solubility Rules                                                    |   |                                                                                                                  | Form a Compound which is |  |
|---------------------------------------------------------------------|---|------------------------------------------------------------------------------------------------------------------|--------------------------|--|
| <b>Negative Ion Plus Positive Ion</b>                               |   |                                                                                                                  |                          |  |
| Any neg. ion + group 1a ions Soluble                                |   |                                                                                                                  |                          |  |
| Any neg. ion + Ammonium ion Soluble                                 |   |                                                                                                                  |                          |  |
| Nitrate + any positive ion Soluble                                  |   |                                                                                                                  |                          |  |
| Acetate                                                             | + | any positive ion except Ag <sup>+</sup> or Hg <sup>2+</sup>                                                      | Soluble                  |  |
| bromide or iodide + Any other positive ion                          | + | Ag <sup>+</sup> , Pb <sup>2+</sup> , Hg <sup>2+</sup> or Cu <sup>+</sup>                                         | Not Soluble              |  |
| Sulfate                                                             | + | Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> , Ra <sup>2+</sup> , Ag <sup>+</sup> or Pb <sup>2+</sup>  | Not Soluble              |  |
|                                                                     | + | Any other positive ion                                                                                           | Soluble                  |  |
| Sulfide                                                             | + | group 1a ions or ammonium                                                                                        | Soluble                  |  |
|                                                                     | + | Be <sup>2+</sup> , Mg <sup>2+</sup> , Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> or Ra <sup>2+</sup> | Soluble                  |  |
|                                                                     | + | Any other positive ion                                                                                           | Not Soluble              |  |
| Hydroxide                                                           | + | group 1a ions or ammonium                                                                                        | Soluble                  |  |
|                                                                     | + | Any other positive ion                                                                                           | Not Soluble              |  |
| Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion | + | group 1a ions or ammonium                                                                                        | Soluble                  |  |
|                                                                     | + | Any other positive ion                                                                                           | Not Soluble              |  |

## Equations

|                                          |                                            |
|------------------------------------------|--------------------------------------------|
| D = m/v                                  | c = specific heat                          |
| Molarity = $\frac{\text{mol}}{\text{V}}$ | d = density                                |
| $P_1 V_1 = P_2 V_2$                      | I = initial amount                         |
| $\frac{T_1}{T_2}$                        | F = final amount                           |
| Q = mc ΔT                                | m = mass                                   |
| nRTPV =                                  | M = molarity                               |
| F = I (1/2) *                            | n = number of moles & number of half lives |
| n = Total Time                           | P = pressure                               |
| Half Life Time                           | R = ideal gas constant                     |
|                                          | Q = heat energy                            |
|                                          | T = temperature                            |
|                                          | V = volume                                 |
|                                          | X = amount of matter left                  |

## Constants

|                                    |                                     |
|------------------------------------|-------------------------------------|
|                                    | 1 atm = 101.3 kPa Pressure          |
|                                    | 1 atm = 760 mm Hg                   |
|                                    | R = 8.314 L x kPa / mol°C Universal |
| Constant or 0.0821 L x atm / mol°C |                                     |
| Avogadro's number                  | 6.02 X                              |
| Specific Heats (J / g°C)           |                                     |
|                                    | Water = 4.180                       |
|                                    | Aluminum = 0.895                    |
|                                    | Copper = 0.387                      |
|                                    | Iron = 0.448                        |

## Activity Series in Decreasing Activity

| Metals    |
|-----------|
| Lithium   |
| Rubidium  |
| Potassium |
| Calcium   |
| Sodium    |
| Magnesium |
| Aluminum  |
| Manganese |
| Zinc      |
| Iron      |
| Nickel    |
| Tin       |
| Lead      |
| Hydrogen  |
| Copper    |
| Silver    |
| Platinum  |
| Gold      |
| Nonmetals |
| Fluorine  |
| Chlorine  |
| Bromine   |
| Iodine    |

## Polyatomic Ions

| Name                              | Formula                                                   |
|-----------------------------------|-----------------------------------------------------------|
| Ammonium                          | NH <sub>4</sub> <sup>+</sup>                              |
| Acetate                           | C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> |
| Carbonate                         | CO <sub>3</sub> <sup>-2</sup>                             |
| Bicarbonate or Hydrogen carbonate | HCO <sub>3</sub> <sup>-</sup>                             |
| Hypochlorite                      | ClO <sup>-</sup>                                          |
| Chlorite                          | ClO <sub>2</sub> <sup>-</sup>                             |
| Chlorate                          | ClO <sub>3</sub> <sup>-</sup>                             |
| Perchlorate                       | ClO <sub>4</sub> <sup>-</sup>                             |
| Hydroxide                         | OH <sup>-</sup>                                           |
| Cyanide                           | CN <sup>-</sup>                                           |
| Nitrate                           | NO <sub>3</sub> <sup>-</sup>                              |
| Nitrite                           | NO <sub>2</sub> <sup>-</sup>                              |
| Chromate                          | CrO <sub>4</sub> <sup>-2</sup>                            |
| Oxalate                           | C <sub>2</sub> O <sub>4</sub> <sup>-2</sup>               |
| Sulfate                           | SO <sub>4</sub> <sup>-2</sup>                             |
| Sulfite                           | SO <sub>3</sub> <sup>-2</sup>                             |
| Hydrogen sulfate                  | HSO <sub>4</sub> <sup>-</sup>                             |
| Phosphate                         | PO <sub>4</sub> <sup>-3</sup>                             |
| Phosphite                         | PO <sub>3</sub> <sup>-3</sup>                             |
| Hydrogen phosphate                | HPO <sub>4</sub> <sup>-2</sup>                            |

## Metric Conversions

| Unit   | Abbr. | Meaning          |
|--------|-------|------------------|
| Mega-  | M     | 10 <sup>6</sup>  |
| kilo-  | k     | 10 <sup>3</sup>  |
| hecto- | h     | 10 <sup>2</sup>  |
| deca-  | da    | 10 <sup>1</sup>  |
| Liter  | L     | 1                |
| gram   | g     | 1                |
| meter  | m     | 1                |
| deci-  | d     | 10 <sup>-1</sup> |
| centi- | c     | 10 <sup>-2</sup> |
| milli- | m     | 10 <sup>-3</sup> |
| micro- | μ     | 10 <sup>-6</sup> |
| nano-  | n     | 10 <sup>-9</sup> |

## Calorie Conversions

|                    |
|--------------------|
| Water = 1 cal/ g°C |
| 1000 cal = 1 Cal   |

# Quality Core Chemistry Formula Sheet

**Jerry A. Nathanson**



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