

# Qualitative Analysis and Chemical Bonding

## Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

## Introduction

- Types of atoms are not equal in chemical bond: **electronegativity**
- In order to understand bond of atoms/molecules we need to know how they are held together, strong
- In order to know a bond between two atoms, strong, weak, we need to know the degree of the attraction

### Electronegativity

- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent compound
- More electronegative atom is more able to attract electrons in a bond, hence, more negative charge, more strong
- Electronegativity increases with the strength of charge
- Periodic table electronegativity decreases **downward**
- From electronegative towards negative electronegativity charge, we can predict the bond **strength**



### Chemical bonding

How atoms/molecules bond together, how they are held together, strong, weak, we need to know the degree of the attraction

### Electronegativity

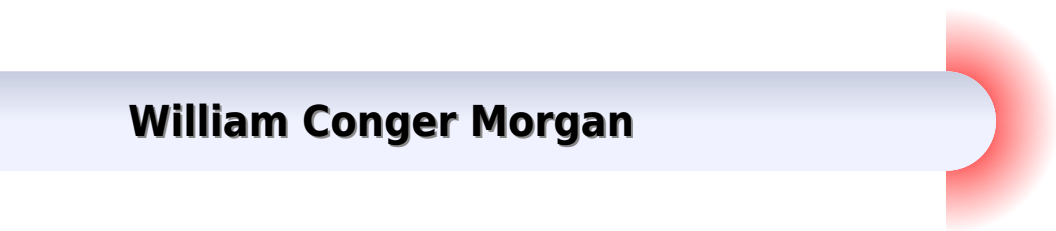


Electronegativity is a measure of the ability of an atom to attract electrons towards itself in a covalent compound. It is a relative property, and the scale is arbitrary. The most electronegative element is fluorine, and the least electronegative element is cesium.

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**William Conger Morgan**



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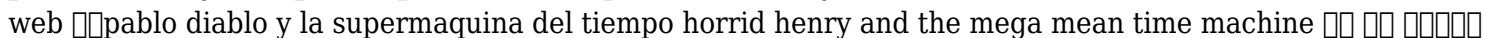
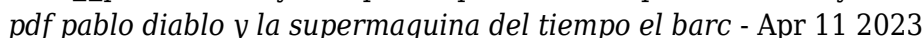
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