

8 • Electron Configuration & the Periodic Table**AP QUESTION**

- (a) Write the ground state electron configuration (short form or long form) for an arsenic atom.

As $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$ or As $[Ar] 4s^2 3d^{10} 4p^3$ or $[Ar] 3d^{10} 4s^2 4p^3$

- (b) Give one permissible set of four quantum numbers for each of the outermost electrons in a single As atom when it is in its ground state.

 $\overset{4p}{\begin{array}{|c|c|c|c|} \hline 4 & 1 & -1 & +\frac{1}{2} \\ \hline \end{array}}$ $\overset{4p}{\begin{array}{|c|c|c|c|} \hline 4 & 1 & 0 & +\frac{1}{2} \\ \hline \end{array}}$ $\overset{4p}{\begin{array}{|c|c|c|c|} \hline 4 & 1 & +1 & +\frac{1}{2} \\ \hline \end{array}}$ $\left(\begin{array}{c} +\frac{1}{2} \\ -\frac{1}{2} \end{array} \right)$ $\overset{4s}{\begin{array}{|c|c|c|c|} \hline 4 & 0 & 0 & +\frac{1}{2} \\ \hline \end{array}}$ $\overset{4s}{\begin{array}{|c|c|c|c|} \hline 4 & 0 & 0 & -\frac{1}{2} \\ \hline \end{array}}$

- (c) Is an isolated arsenic atom in the ground state
- paramagnetic
- or diamagnetic? Explain briefly.

Paramagnetic $4p^3 = \uparrow \uparrow \uparrow$ half-filled orbitals = paramagnetic.

- (d) What ion would As be expected to form?
- As³⁻
-
- Write the electron configuration for the arsenic ion.

As $[Ar] 4s^2 3d^{10} 4p^6$ or As $[Ar] 3d^{10} 4s^2 4p^6$ As gains electrons and has the same electron configuration as Kr.

3-	2-	1-	He
N	O	F	Ne
P	S	Cl	Ar
As	Se	Br	Kr

Periodic Worksheet Chemistry If8766 Answer

Rob King

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Periodic Worksheet Chemistry If8766 Answer

1. Understanding the eBook Periodic Worksheet Chemistry If8766 Answer
 - The Rise of Digital Reading Periodic Worksheet Chemistry If8766 Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Worksheet Chemistry If8766 Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Worksheet Chemistry If8766 Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Worksheet Chemistry If8766 Answer
 - Personalized Recommendations
 - Periodic Worksheet Chemistry If8766 Answer User Reviews and Ratings
 - Periodic Worksheet Chemistry If8766 Answer and Bestseller Lists
5. Accessing Periodic Worksheet Chemistry If8766 Answer Free and Paid eBooks
 - Periodic Worksheet Chemistry If8766 Answer Public Domain eBooks
 - Periodic Worksheet Chemistry If8766 Answer eBook Subscription Services
 - Periodic Worksheet Chemistry If8766 Answer Budget-Friendly Options
6. Navigating Periodic Worksheet Chemistry If8766 Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Worksheet Chemistry If8766 Answer Compatibility with Devices
 - Periodic Worksheet Chemistry If8766 Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Worksheet Chemistry If8766 Answer
 - Highlighting and Note-Taking Periodic Worksheet Chemistry If8766 Answer
 - Interactive Elements Periodic Worksheet Chemistry If8766 Answer
8. Staying Engaged with Periodic Worksheet Chemistry If8766 Answer

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Periodic Worksheet Chemistry If8766 Answer
- 9. Balancing eBooks and Physical Books Periodic Worksheet Chemistry If8766 Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Worksheet Chemistry If8766 Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Worksheet Chemistry If8766 Answer
 - Setting Reading Goals Periodic Worksheet Chemistry If8766 Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Worksheet Chemistry If8766 Answer
 - Fact-Checking eBook Content of Periodic Worksheet Chemistry If8766 Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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