#### Period Date Tuesday, November 1 2016

## Periodic Trends

#### ATOMIC RADIUS

- What trend in atomic radius do you see as you go down a group/family on the periodic table?
- What causes this trend?
- 3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
- 4. What causes this trend?
- Circle the atom in each pair that has the largest atomic radius.
  - a) Al B

- b) S O
- c) Br Cl

- d) Na Al e) O F f) Mo Ca
- Put the following elements in order from smallest to largest atomic radius and explain why: C. O. Sn. Sr.

#### ELECTRONEGATIVITY

- Define electronegativity
- 8. How does the ionic radius of a nonmetal compare with its atomic radius?
- 9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
- 10. What causes this trend?
- 11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
- 12. What causes this trend?
- Circle the atom in each pair that has the greater electronegativity.
- a) Ca Ga

- b) Li O c) Cl S d) Br As
- e) Ba Sr

#### GENERAL QUESTIONS

- 14. Which group tends to form +1 ions?
- 15. Which group tends to form +2 ions?
- 16. Which group tends to form -1 ions?
- 17. Which group tends not to form ions or react?

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#### **Table of Contents Periodic Trends Worksheet Extension Questions Answers**

- 1. Understanding the eBook Periodic Trends Worksheet Extension Questions Answers
  - The Rise of Digital Reading Periodic Trends Worksheet Extension Questions Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Worksheet Extension Questions Answers
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Worksheet Extension Questions Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Worksheet Extension Questions Answers
  - Personalized Recommendations

- Periodic Trends Worksheet Extension Questions Answers User Reviews and Ratings
- Periodic Trends Worksheet Extension Ouestions Answers and Bestseller Lists
- 5. Accessing Periodic Trends Worksheet Extension Questions Answers Free and Paid eBooks
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  - Periodic Trends Worksheet Extension Questions Answers eBook Subscription Services
  - Periodic Trends Worksheet Extension Questions Answers Budget-Friendly Options
- 6. Navigating Periodic Trends Worksheet Extension Questions Answers eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Periodic Trends Worksheet Extension Questions Answers Compatibility with Devices
  - Periodic Trends Worksheet Extension Questions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Worksheet Extension Questions Answers
  - Highlighting and Note-Taking Periodic Trends Worksheet Extension Questions Answers
  - Interactive Elements Periodic Trends Worksheet Extension Ouestions Answers
- 8. Staying Engaged with Periodic Trends Worksheet Extension Questions Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Worksheet Extension Questions Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Worksheet Extension Questions Answers
  - ∘ Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Periodic Trends Worksheet Extension Questions Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Worksheet Extension Questions Answers
  - Setting Reading Goals Periodic Trends Worksheet Extension Questions Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Worksheet Extension Questions Answers
  - Fact-Checking eBook Content of Periodic Trends Worksheet Extension Questions Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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