

Fill in the gaps in statements and/or equations.

1. The theory that explains the behaviour of gases on the molecular level is called the kinetic theory, which is based on assumptions related to the ideal gas of non-interacting particles.
2. Use the kinetic theory to explain why a balloon filled with air expands when it is brought inside a warmer room from the cold outdoors.
 According to the kinetic theory, particles have more kinetic energy. Consequently, the pressure is higher in warmer rooms. Within the container, which has a fixed volume and number of particles, the pressure is higher (fixed volume). Therefore, the volume of the gas must increase to maintain the pressure in the room. The volume of the gas is proportional to the absolute temperature.
3. Define "1 bar" and "1 Pa".
 1 bar is defined as 10⁵ Pa and 1 Pa is defined as 1 N/m².
4. Gases deviate most from ideal behaviour under conditions of high pressure and low temperature.
5. Pressure is defined as force per area. The earth's atmosphere has weight, which creates atmospheric pressure.
6. Standard temperature has been established as 0 °C or 273 K. Standard pressure has been established as 1013 hPa, 1013 mbar, 1013 hPa, or 1013 hPa. The standard unit for standard temperature and pressure is 0 °C.
7. A manometer is used to measure atmospheric pressure, which is atmospheric pressure or atmospheric pressure.
8. With atmospheric pressure increasing, how does the height of the column of mercury change? It increases.
9. At constant pressure, the volume of a sample of gas increases as temperature increases.
10. At constant pressure, the volume of a sample of gas is proportional to the absolute temperature.

Review Sheet Unit 9 Chemistry Answers

Dixie Goss



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