

Give the element and the electron configuration notation for the following:

Quantum Number	Element Identification	Noble Gas Notation
2, 1, +1, $+\frac{1}{2}$	N	[He] 2s² 2p³
3, 2, +1, $-\frac{1}{2}$	Cu	[Ar] 4s² 3d⁹
5, 3, 0, $-\frac{1}{2}$	Es	[Rn] 7s² 5f¹¹
7, 0, 0, $-\frac{1}{2}$	Ra	[Rn] 7s²
6, 1, -1, $-\frac{1}{2}$	Po	[Xe] 6s² 4f¹⁴ 5d¹⁰ 6p⁴
6, 0, 0, $+\frac{1}{2}$	Cs	[Xe] 6s¹
5, 1, +1, $-\frac{1}{2}$	Xe	[Kr] 5s² 4d¹⁰ 5p⁶
5, 2, -1, $-\frac{1}{2}$	Ir	[Xe] 6s² 4f¹⁴ 5d⁷
4, 2, -2, $+\frac{1}{2}$	Y	[Kr] 5s² 4d¹
1, 0, 0, $-\frac{1}{2}$	He	1s²

Given the element and quantum number:

Noble-Gas Electron Configuration	Element Identification	Quantum Number
[Rn] 7s ² 5f ⁴	U	5, 3, +1, $+\frac{1}{2}$
[Ar] 4s ² 3d ¹⁰ 4p ²	Ge	4, 1, 0, $+\frac{1}{2}$
[Ar] 4s ² 3d ³	V	3, 2, 0, $+\frac{1}{2}$
[He] 2s ¹	Li	2, 0, 0, $+\frac{1}{2}$
[Kr] 5s ² 4d ³	Nb	4, 2, 0, $+\frac{1}{2}$

Quantum Numbers Practice Problems With Answers

David Kirk



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