

• *Electrons and Quantum Study Guide*

(36 pts) MC = 24 pts choose between two = 1 pt Short answer = 7 pts Matching = 4

(Regular chemistry students may use a 3x5 in. card with notes on it for the test. Something to include would be the π -filling chart, etc. [If your card is larger than 3x5 in., I will cut it to size.] (Phonons will be provided with the π -filling chart only.)

Be able to:

- ☐ Identify and apply the Aufbau principle, Hund's rule, and Pauli Exclusion Principle
- ☐ Identify and write proper orbital notations, electron configurations, and noble-gas notations of elements up to atomic # 86.
- ☐ Identify what a photon is and when it is emitted from an atom.
- ☐ Distinguish between an emission and absorption spectra and how and why they are produced.
- ☐ Know what each of the 4 quantum numbers stand for.
- ☐ Recall the 4 orbital shapes and the letter they correspond to
- ☐ Use the electron filling order chart
- ☐ Correlate period numbers to energy levels and the "A" group numbers to valence electrons.
- ☐ Explain the emission spectrum of hydrogen using the Bohr model.
- ☐ Know the descriptions of the Heisenberg Uncertainty Principle, Schrödinger wave equation, atomic orbital, and the quantum mechanical model of the atom.
- ☐ Express the electron filling order exceptions of copper and chromium.
- ☐ Write from memory the numerical value of the speed of all electromagnetic radiation.
- ☐ Express the visible light spectrum and the electromagnetic spectrum in sequence of wavelength, frequency, or energy.
- ☐ Identify the contributions of Planck, Einstein, de Broglie, Schrodinger, Heisenberg, and Bohr
- ☐ Tell how many orbitals & electrons can exist in a given energy level.
- ☐ Tell how many orbitals the electrons occupy in the ground state of a given element.
- ☐ Tell the number of orientations of a given sublevel.
- ☐ Tell the number of orbitals that inner-shell (level) electrons occupy.
- ☐ Know the terms wavelength, frequency, and amplitude.
- ☐ Be able to adequately compare and contrast the Bohr and Quantum model.

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Didier Musso



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