Quantum Numbers Practice Worksheet

Name: KEY

Give the element and the electron configuration notation for the following:

Quantum Number Identifica Noble G		Noble Gas Notation	
2, 1, +1, +1/2	N	[He] 2s ² 2p ³	
3, 2, +1, -1/2	Cu	[Ar] 4s ² 3d ⁹	
5, 3, 0, -1/2	Es	[Rn] 7s ² 5f ¹¹	
7, 0, 0, -1/2	Ra	[Rn] 7s ²	
6, 1, -1, -1/2	Po	[Xe] 6s ² 4f ¹⁴ 5d ¹⁰ 6p ⁴	
6, 0, 0, +1/2	Cs	[Xe] 6s1	
5, 1, +1, -1/2	Xe	[Kr] 5s ² 4d ¹⁰ 5p ⁶	
5, 2, -1, -1/2	lr	[Xe] 6s ² 4f ¹⁴ 5d ⁷	
4, 2, -2, +1/2	Y	[Kr] 5s ² 4d ¹	
1, 0, 0, -1/2	He	1s²	

Given the element and quantum number:

Noble-Gas Electron Configuration	Element Identification	Quantum Number
[Rn] 7s ² 5f ⁴	U	5, 3, +1, +1
[Ar] 4s ² 3d ¹⁰ 4p ²	Ge	4, 1, 0, +1
[Ar] 4s² 3d ³	v	3, 2, 0, +1
[He] 2s ¹	Li	2, 0, 0, +1
[Kr] 5s ² 4d ³	Nb	4, 2, 0,+

Quantum Numbers Practice

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Southam, Gwendolyn Lawrie, Natalie
Williamson, Christopher Thompson

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ionic bonding ionic compounds Covalent Bonds and Formulas nomenclature electronic molecular geometries octet rule polar molecules The Mole Concept formula stoichiometry Chemical Reactions balancing equations reaction types Stoichiometry limiting reactant equations yields heat of reaction The Behavior of Gases molecular structure properties combined gas law universal gas law Condensed Phases Solids and Liquids intermolecular forces of attraction phase change phase diagrams Solutions and Their Behavior concentration solubility colligate properties dissociation ions in solution Chemical Kinetics reaction rates factors that affect rates Chemical Equilibrium forward reverse reaction rates equilibrium constant Le Chatelier s principle solubility product constant Acids Bases strong weak acids and bases hydrolysis of salts pHNeutralization dissociation of water acid base indicators acid base titration buffers Thermochemistry bond breaking formation heat of reaction formation Hess law entropy Gibb s free energy Electrochemistry oxidation reduction electrochemical cells Nuclear Chemistry radioactivity nuclear equations nuclear energy Organic Chemistry straight chain aromatic hydrocarbons functional groups Chemistry Glossary Energy Technology and Directions for the Future John R. Fanchi, 2013-10-22 Energy Technology and Directions for the Future presents the fundamentals of energy for scientists and engineers It is a survey of energy sources that will be available for use in the 21st century energy mix The reader will learn about the history and science of several energy sources as well as the technology and social significance of energy Themes in the book include thermodynamics electricity distribution geothermal energy fossil fuels solar energy nuclear energy alternate energy wind water biomass energy and society energy and the environment sustainable development the hydrogen economy and energy forecasting The approach is designed to present an intellectually rich and interesting text that is also practical This is accomplished by introducing basic concepts in the context of energy technologies and where appropriate in historical context Scientific concepts are used to solve concrete engineering problems The technical level of presentation presumes that readers have completed college level physics with calculus and mathematics through calculus of several variables The selection of topics is designed to provide the reader with an introduction to the language concepts and techniques used in all major energy components that are expected to contribute to the 21st century energy mix Future energy professionals will need to understand the origin and interactions of these energy components to thrive in an energy industry that is evolving from an industry dominated by fossil fuels to an industry working with many energy sources Presents the fundamentals of energy production for engineers scientists engineering professors students and anyone in the field who needs a technical discussion of energy topics Provides engineers with a valuable expanded knowledge base using the U S National Academy of Sciences content standards Examines the energy options for the twenty first century as older energy sources quickly become depleted Quantities, Units and Symbols in Physical Chemistry E Richard Cohen, Tom Cvitas, Jeremy G Frey, Bertil Holström, Kozo Kuchitsu, Roberto Marquardt, Ian Mills, Franco Pavese, Martin Quack, Juergen Stohner, Herbert L Strauss, Michio Takami, Anders J Thor, 2007-10-31 The first IUPAC Manual of Symbols and Terminology for Physicochemical

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