

More Practice

1. List the molar ratios you can derive from this balanced chemical equation:



2. List the molar ratios you can derive from this balanced chemical equation:



3. Given the following balanced chemical equation,



how many moles of NaCl can be formed if 3.77 mol of NaOH were to react?

4. Given the following balanced chemical equation,



how many moles of H_2O can be formed if 0.0652 mol of C_5H_{12} were to react?

5. Balance the following unbalanced equation and determine how many moles of H_2O are produced when 1.65 mol of NH_3 react.



6. Trinitrotoluene [$\text{C}_6\text{H}_2(\text{NO}_2)_3\text{CH}_3$], also known as TNT, is formed by reacting nitric acid (HNO_3) with toluene ($\text{C}_6\text{H}_5\text{CH}_3$):



Balance the equation and determine how many moles of TNT are produced when 4.903 mol of HNO_3 react.

7. Chemical reactions are balanced in terms of molecules and in terms of moles. Are they balanced in terms of dozens? Defend your answer.
8. Explain how a chemical reaction balanced in terms of moles satisfies the law of conservation of matter.

Quantities In Chemical Reactions Section Review Answers

Marco Cascella



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