

Qualitative Analysis and Chemical Bonding

Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

Introduction

- Types of atoms are not equal in chemical bonds: **electronegativity**
- In order to understand how all substances behave we need to know how they are bonded, strong
- In order to know a substance behaves like ionic, strong, must first know the nature of the interaction

Electronegativity

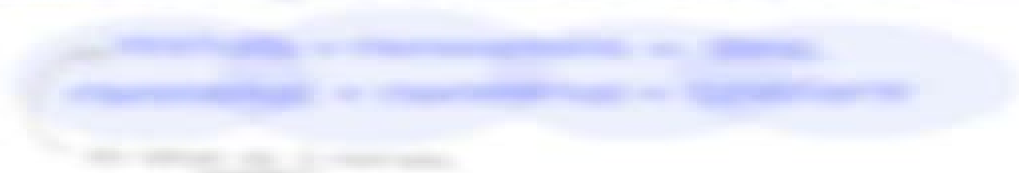
- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent compound
- More electronegative atoms are more likely to attract electrons in a bond, pulling electron density towards them
- Electronegativity increases with increasing atomic number
- Periodic trend: electronegativity increases **diagonally upwards**
- From electronegative fluorine (most electronegative) through noble gas argon to electropositive cesium



Covalent bonding

Two electronegative atoms sharing their outer, or valence, electrons to form a bond.

Molecule is group of atoms held together by covalent bonds.



Electron density is higher in the region between the two atoms, creating a bond. The electron density is higher in the region between the two atoms, creating a bond.

Electron density is higher in the region between the two atoms, creating a bond. The electron density is higher in the region between the two atoms, creating a bond.

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