

Qualitative Analysis and Chemical Bonding

Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

Introduction

- Types of atoms are not equal in chemical bonds: **electronegativity**
- In order to understand how all substances behave we need to know how they are bonded, strong
- In order to know a substance behaves like ionic, strong, must be able to separate its ions completely

Electronegativity

- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent compound
- More electronegative atom is more likely to attract electrons in a bond, pulling electron density towards itself
- Electronegativity increases with the strength of a bond
- More electronegative atoms attract electrons more strongly than less electronegative atoms
- In a bond between two atoms, the more electronegative atom will pull electrons towards itself

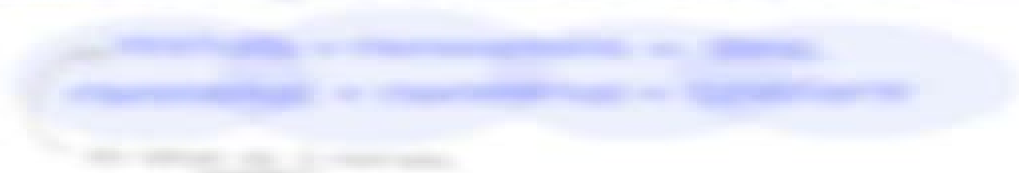


Covalent bonding

Two atoms share electrons, resulting from pairs of electrons from two separate valence shells.

Electronegativity

is a group of atoms that together are similar to an atom.



Electronegativity is a measure of an atom's ability to attract electrons towards itself in a covalent bond. It is a relative property, with fluorine being the most electronegative element.

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Qualitative Analysis And Chemical Bonding Teacher Notes

S Ashworth



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