

Name: _____ Date: _____ Period: _____

Periodic Trends Practice Worksheet

1. What is the difference between electronegativity and ionization energy?
2. Why does fluorine have a higher ionization energy than iodine?
3. Why do elements in the same family generally have similar properties?
4. Which is the largest atom in Group 4a (14)?
5. Which is the smallest atom in Group 7a (17)?
6. Which is the smallest atom in period 5?
7. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
8. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
9. Arrange the following atoms in order of decreasing atomic radius. Na Al P Cl Mg
10. For each of the following pairs, circle the element that is larger?
a) N^{3-} or F^{-} b) Mg^{2+} or Ca^{2+} c) Fe^{2+} or Fe^{3+}
11. Circle the element in each pair has the larger radius?
a) Mg or Mg^{2+} b) O or O^{2-} c) K^{+} or Cl^{-} d) P^{3-} or S^{2-}
12. In each of the following pairs, circle the species with the higher first ionization energy:
a) Li or Cs (b) Cl or Ar (c) Ca or Br (d) Na or Ne (e) B or Be
13. In each of the following pairs, circle the species with the larger atomic radius:
a) Mg or Ba (b) S or S^{2-} (c) Cu^{+2} or Cu (d) He or H (e) Na or Cl
14. Circle the best choice in each list:
a) highest first ionization energy: C, N, Si
b) highest electronegativity: As, Sn, S
c) largest radius: S^{2-} , Cl^{-} , Cl
15. Order the following groups from largest to smallest radii.
a) Ar, Cl, K, S^{2-}
b) C, Al, F, Si
c) Na, Mg, Ar, P
d) I^{-} , Ba^{2+} , Cs^{+1} , F^{-1}

Periodic Trends Extension Questions Worksheet

JG Myers



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Table of Contents Periodic Trends Extensoin Questions Worksheet

1. Understanding the eBook Periodic Trends Extensoin Questions Worksheet
 - The Rise of Digital Reading Periodic Trends Extensoin Questions Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Extensoin Questions Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Extensoin Questions Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Extensoin Questions Worksheet
 - Personalized Recommendations
 - Periodic Trends Extensoin Questions Worksheet User Reviews and Ratings

- Periodic Trends Extensoin Questions Worksheet and Bestseller Lists
- 5. Accessing Periodic Trends Extensoin Questions Worksheet Free and Paid eBooks
 - Periodic Trends Extensoin Questions Worksheet Public Domain eBooks
 - Periodic Trends Extensoin Questions Worksheet eBook Subscription Services
 - Periodic Trends Extensoin Questions Worksheet Budget-Friendly Options
- 6. Navigating Periodic Trends Extensoin Questions Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Extensoin Questions Worksheet Compatibility with Devices
 - Periodic Trends Extensoin Questions Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Extensoin Questions Worksheet
 - Highlighting and Note-Taking Periodic Trends Extensoin Questions Worksheet
 - Interactive Elements Periodic Trends Extensoin Questions Worksheet
- 8. Staying Engaged with Periodic Trends Extensoin Questions Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Extensoin Questions Worksheet
- 9. Balancing eBooks and Physical Books Periodic Trends Extensoin Questions Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Extensoin Questions Worksheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Extensoin Questions Worksheet
 - Setting Reading Goals Periodic Trends Extensoin Questions Worksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Extensoin Questions Worksheet
 - Fact-Checking eBook Content of Periodic Trends Extensoin Questions Worksheet
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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