

## The Periodic Table

### The Structure of the Periodic Table

- The periodic table displays the elements in order of increasing atomic number i.e. in order of increasing number of protons.
- Elements are arranged in 18 vertical columns called groups that are numbered 1 – 18.
- The seven horizontal rows in periodic table are called periods. The periods are numbered 1-7. The first period contains only two elements, hydrogen and helium.
- Elements within the same groups have similar chemical properties because they have the same number of valence electrons. The number of valence electrons is the same as the group number (in Roman numerals).

|                  |   |                  |              |              |              |              |              |              |              |              |              |              |              |              |              |              |             |              |              |             |             |              |              |
|------------------|---|------------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|--------------|-------------|--------------|--------------|-------------|-------------|--------------|--------------|
|                  |   | GROUPS<br>1 - 18 |              |              |              |              |              |              |              |              |              |              |              |              |              |              |             |              |              |             |             |              |              |
| PERIODS<br>1 - 7 | 1 |                  |              |              |              |              |              |              |              |              |              |              |              |              |              |              |             | 18           |              |             |             |              |              |
|                  | 1 | H<br>1.0079      |              |              |              |              |              |              |              |              |              |              |              |              |              |              |             |              | He<br>4.0026 |             |             |              |              |
|                  | 2 | Li<br>6.941      | Be<br>9.0122 |              |              |              |              |              |              |              |              |              |              |              |              |              |             | B<br>10.811  | C<br>12.011  | N<br>14.007 | O<br>15.999 | F<br>18.998  | Ne<br>20.18  |
|                  | 3 | Na<br>22.99      | Mg<br>24.305 |              |              |              |              |              |              |              |              |              |              |              |              |              |             | Al<br>26.982 | Si<br>28.086 | P<br>30.974 | S<br>32.06  | Cl<br>35.453 | Ar<br>39.948 |
|                  | 4 | K<br>39.098      | Ca<br>40.078 | Sc<br>44.956 | Ti<br>47.88  | V<br>50.942  | Cr<br>51.996 | Mn<br>54.938 | Fe<br>55.847 | Co<br>58.933 | Ni<br>58.693 | Cu<br>63.546 | Zn<br>65.39  | Ga<br>69.723 | Ge<br>72.64  | As<br>74.922 | Se<br>78.96 | Br<br>79.904 | Kr<br>83.8   |             |             |              |              |
|                  | 5 | Rb<br>85.468     | Sr<br>87.62  | Y<br>88.906  | Zr<br>91.224 | Nb<br>92.906 | Mo<br>95.94  | Tc<br>97.907 | Ru<br>101.07 | Rh<br>102.91 | Pd<br>106.42 | Ag<br>107.87 | Cd<br>112.41 | In<br>114.82 | Sn<br>118.71 | Sb<br>121.76 | Te<br>127.6 | I<br>126.9   | Xe<br>131.29 |             |             |              |              |
|                  | 6 | Cs<br>132.91     | Ba<br>137.33 | La<br>138.91 | Hf<br>178.49 | Ta<br>180.95 | W<br>183.84  | Re<br>186.21 | Os<br>190.23 | Ir<br>192.22 | Pt<br>195.08 | Au<br>196.97 | Hg<br>200.59 | Tl<br>204.38 | Pb<br>207.2  | Bi<br>208.98 | Po<br>209   | At<br>210    | Rn<br>222    |             |             |              |              |
|                  | 7 | Fr<br>223        | Ra<br>226    | Ac<br>227    | Rf<br>261    | Ha<br>262    | Sg<br>266    | Nh<br>271    | Hs<br>277    | Mt<br>288    | Ds<br>289    |              |              |              |              |              |             |              |              |             |             |              |              |

### Energy levels and period numbers

The electrons of an atom are arranged in energy levels. The period number tells us in which energy level an atom has its valence electrons.

For example:

| Element     | In which Group? | In which Period? | How many valence electrons | Which energy level are valence electrons? |
|-------------|-----------------|------------------|----------------------------|-------------------------------------------|
| Phosphorous | 15              | 3                | 5                          | 3 <sup>rd</sup> energy level              |
| Potassium   |                 |                  |                            |                                           |
| Calcium     |                 |                  |                            |                                           |
| Oxygen      |                 |                  |                            |                                           |

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**K Payea**



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