

PreAP Chemistry Review: Solutions 1

1. Explain what is meant by the expression "like dissolves like".
2. An unknown compound is observed to mix with benzene (a nonpolar solvent) but not with water. Is the unknown compound ionic or covalent?
3. If the unknown compound from #2 is a liquid, will it be able to dissolve table salt?
4. Explain # 2 & 3
5. What are the chemical characteristics of a good dry-cleaning solvent?
6. Some industrial plants use water from nearby rivers and streams as a coolant. When the water is returned to the river or stream, the water is warmer than it was originally. This is referred to as "thermal pollution". Using your knowledge of solubility, why might this thermal pollution be harmful to fish?
7. After a bottle of carbonated drink has been open for a while, it tastes "flat". Explain why.
8. For most solid solutes, the degree of solubility in a liquid solvent (increases, decreases) with an increase in the temperature of the solvent.
9. Describe what happens to the degree of solubility of a gaseous solute in a liquid:
 - a) with a decrease in the temperature of the solvent.
 - b) with an increase in pressure
10. Suppose you prepare a 0.200 M salt solution in a 250-mL volumetric flask and then accidentally spill some of it. What happens to the concentration of the solution in the bottle?
11. Calculate the number of grams of each solute that has to be used to make each of the following solution.
 - (a) 250 mL of 0.100 M NaCl
 - (b) 100 mL of 0.440 M $C_6H_{12}O_6$ (glucose)
 - (c) 500 mL of 0.500 M H_2O_2
12. The following questions refer to a 0.20 M solution of $BaCl_2$. The molecular mass of $BaCl_2$ is 208.23 g/mole.
 - a) Calculate the moles of $BaCl_2$ in 2.00 mL of solution.
 - b) Calculate the mg and grams of $BaCl_2$ in 2.00 mL of the solution.
 - c) Calculate the volume in mL that contains 5.0 moles of $BaCl_2$.
 - d) Calculate the volume in mL that contains 41.6 mg of $BaCl_2$.
12. What is the molality of a solution containing 7.8 g of $MgCl_2$ in 725 g of water?
13. What is the difference between a strong electrolyte and a weak electrolyte?
14. Compare and contrast dissociation and ionization.
16. Predict whether the following substances will be soluble or insoluble when mixed.
 - a. water and methanol (CH_3OH)
 - b. wax and CCl_4
17. You have a flask containing a solution of NaCl. How could you determine whether the solution is saturated or unsaturated?

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WJ Hussar



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