Gas Laws Worksheet

atm = 760.0 mm Hg = 101.3 kPa= 760.0 torr

Boyle's Law Problems:

PV. TRV

 If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(749-4)(22.51) = (725-4) V_2$$

 $V_2 = \frac{(749-4)(22.51)}{(725-4)}$

V2 - 23.2L

A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L.
 What is the pressure in the container if the temperature remains constant?

 What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

 A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

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