



Ionisation - when a particle has enough energy to knock an electron out of an atom.

Ionisation energy - the minimum energy required to remove an electron from an atom.

Ionisation potential - the minimum voltage required to remove an electron from an atom.

Photoelectric effect - when a photon hits an electron and it is ejected from the atom. The photon must have enough energy to overcome the work function of the metal.

Work function - the minimum energy required to remove an electron from a metal.

Radioactivity - the process by which an unstable nucleus loses energy by emitting radiation.

Alpha particles - consist of two protons and two neutrons, and are highly ionising.

Beta particles - consist of a single electron or positron, and are less ionising than alpha particles.

Gamma rays - consist of high-energy photons, and are not ionising.

Half-life - the time taken for half of a sample of a radioactive substance to decay.

Decay constant - a measure of the probability that a nucleus will decay per unit time.

Activity - the number of decays per second in a sample of a radioactive substance.

Ionising radiation - radiation that has enough energy to ionise atoms.

Non-ionising radiation - radiation that does not have enough energy to ionise atoms.

Quantum mechanics - the study of the behavior of matter and energy at the atomic and subatomic level.

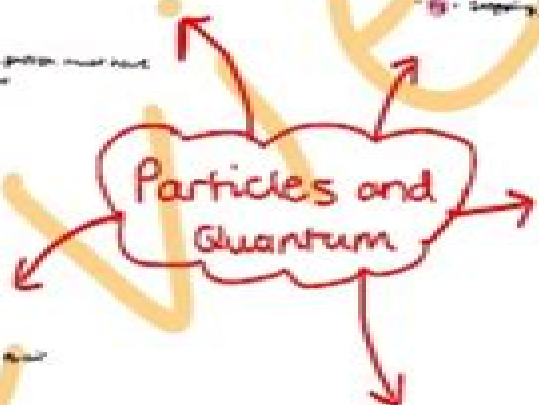
Wave-particle duality - the concept that particles can exhibit both wave-like and particle-like properties.

Heisenberg uncertainty principle - the principle that it is impossible to know both the position and momentum of a particle with perfect accuracy.

Particle	Mass (kg)	Charge (C)	Location
Proton	1.67×10^{-27}	$+1.6 \times 10^{-19}$	Nucleus
Neutron	1.67×10^{-27}	0	Nucleus
Electron	9.1×10^{-31}	-1.6×10^{-19}	Electron cloud

Charge, spin and baryon number must always be conserved.

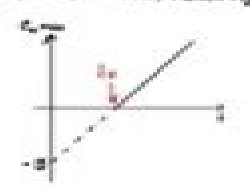
Strong force is only exerted in very close proximity.



Photoelectric effect

the process of light energy ionising electrons in the surface of a metal, liberating them.

- $E_{\text{photon}} = hf$
- $E_{\text{photon}} = \phi + E_{\text{kin}}$ (energy of photon)
- ϕ - work function - minimum energy to liberate electrons
- E_{kin} - kinetic energy of electron after it is liberated
- f - frequency
- h - Planck's constant - minimum frequency needed for electrons to be liberated



The photoelectric effect supports the particle nature of light due to the discrete threshold frequency and maximum kinetic energy. It also shows that the intensity of light does not affect the kinetic energy of the electrons, but it does affect the number of electrons emitted per second.

Stopping potential - the potential needed to stop any electrons from reaching the other side.

$E_{\text{photon}} = \phi + eV_s$

e - charge of electron

V_s - stopping potential

de Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle was proven by diffraction experiments. This is because the electron can diffract, which is a wave property.

$\lambda = \frac{h}{p}$

λ - de Broglie wavelength

p - momentum / mass x velocity

Electron diffraction - the process by which electrons are diffracted by a crystal lattice.

Davisson-Germer experiment - the experiment that demonstrated the wave nature of electrons.

Fluorescent tubes



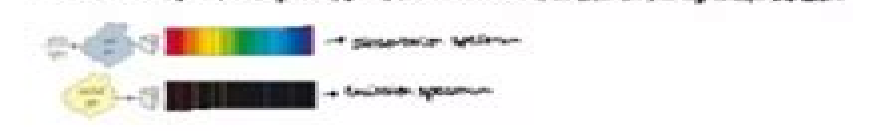
These tubes contain mercury, under a high voltage is applied. This excites the mercury atoms, which then emit light. The light is then converted to visible light by the phosphor coating.

When the gas is excited, it emits light. The light is then converted to visible light by the phosphor coating.

Energy levels

Electrons can be excited to higher energy levels. When they fall back to their ground state, they emit light. The energy of the light is equal to the difference in energy between the two levels.

Excitation energy - the energy required to move an electron from its ground state to an excited state.



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Table of Contents Physics Aqa Empa 2paper

1. Understanding the eBook Physics Aqa Empa 2paper
 - The Rise of Digital Reading Physics Aqa Empa 2paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Aqa Empa 2paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics Aqa Empa 2paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Aqa Empa 2paper
 - Personalized Recommendations
 - Physics Aqa Empa 2paper User Reviews and Ratings

- Physics Aqa Empa 2paper and Bestseller Lists
- 5. Accessing Physics Aqa Empa 2paper Free and Paid eBooks
 - Physics Aqa Empa 2paper Public Domain eBooks
 - Physics Aqa Empa 2paper eBook Subscription Services
 - Physics Aqa Empa 2paper Budget-Friendly Options
- 6. Navigating Physics Aqa Empa 2paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Physics Aqa Empa 2paper Compatibility with Devices
 - Physics Aqa Empa 2paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics Aqa Empa 2paper
 - Highlighting and Note-Taking Physics Aqa Empa 2paper
 - Interactive Elements Physics Aqa Empa 2paper
- 8. Staying Engaged with Physics Aqa Empa 2paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Aqa Empa 2paper
- 9. Balancing eBooks and Physical Books Physics Aqa Empa 2paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Aqa Empa 2paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Aqa Empa 2paper
 - Setting Reading Goals Physics Aqa Empa 2paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Aqa Empa 2paper
 - Fact-Checking eBook Content of Physics Aqa Empa 2paper
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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