(2)

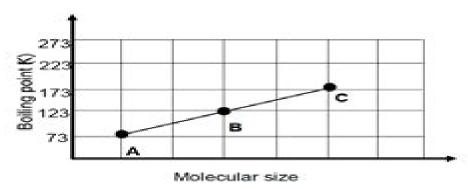
(1)

(3) [13]

QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The **letters A**, **B** and **C** represent the compounds CH₄, C₄H₆ and C₃H₈ respectively.



- 2.1.1 Define the term boiling point.
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (A, B or C) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph.
 (2)
- 2.2 Consider the two molecules in the table below.

| NAME OFSUBSTANCE | FORMULA | MOLACULAR MASS (g) | BOILING POINT |
|---------------------|-----------------|-----------------------|---------------|
| Ammonia | NH ₃ | 17 | - 33 |
| Phosphine | PH ₃ | × | - 87,4 |

- 2.2.1 Determine the molecular mass (x) of phosphine
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES.

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Rachel Sandford

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