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GALAXY ACE 2

Samsung Ace 2 Cell Phones Accessory Owners Manual

M Carnoy



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two substances c_a and c_b the mole fractions of each would be written as

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water $n_a = 54 \text{ g} / 18 \text{ g mol}^{-1} = 3 \text{ mol}$ total number of moles $n_a + n_b = 0.5 + 3 = 3.5 \text{ mol}$ mole fraction of solute ethyl alcohol $x_b = n_b / n_a + n_b$

$= 0.5 / 3.5 = 0.1429$

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40 g of H_2O molar mass of H_2O is 18 gm mole and molar mass of CH_3OH is 32 gm mole solution moles of CH_3OH $4.5 / 32 = 0.14$

mole moles of H_2O $40 / 18 = 2.2$ moles therefore according to the equation

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time the total pressure $p_a = x_a p_t$ mole fraction $x_a = \frac{\text{moles } a}{\text{moles total}}$ right sum $n_1 + \dots + x_n = 1$ mole fraction of substance

a is often denoted by x_a and

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by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction

mole fraction is not dependent on the temperature

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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