Specific Heat

1. A 30g piece of metal absorbs 1,200 joules of heat energy and its temperature changes from 20°C to 175°C. Calculate the specific capacity of the metal. What is the likely metal?

2. How many joules of heat are needed to raise the temperature of 15.0 g of iron from 11°C to 68°C, if the specific heat of iron is 0.45 J/g°C?

3. To what temperature will a 50.0 g piece of brass raise if it absorbs 6 kilojoules of heat and its specific heat capacity is 0.38 J/g°C? The initial temperature of the brass is 20.0°C.

4. Identify the substance if 1500.0 g of the substance absorbs 74,500 joules of heat, and its temperature changes from 25°C to 40°C.

5. 100.0 g of 1.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.

6. 25.0 g of antimony is heated from 10°C to 180°C, and absorbs 750 joules of heat in the process. Calculate the specific heat capacity of antimony.

7. What is the specific heat capacity of quartz glass if 200.0 g of the glass absorbs 2500 joules of heat and the temperature rises 18.5°C?

8. If a sample of titanium is initially at 25°C, what is its final temperature if 250.0 g of titanium absorbs 1.0 kilojoule of heat, and the specific heat of titanium is 0.523 J/g°C?

9. How much energy must be absorbed by 300.0 g of steel to increase its temperature from 100.0°C to 450.0°C? (C of steel = 49.0J/g °C)

10. When 55.0 g of platinum drops in temperature from 300.0°C to 250.0°C, how much heat energy is released? (C of platinum = 130 J/g °C)

11. How much energy is required to heat 125.0 g of lithium from 2.0°C to 30.0°C? (C of lithium = 3.6 J/g°C)

12. How much heat (in kJ) is given out when 350.0 g of lead cools from 200.0°C to 20.0°C?

Practice Problems For Specific Heat

Nuggenhalli S. Nandagopal, PE

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