

#### Aqueous Equilibria: Part 1 - Buffers & Titrations

Name

#### Multiple Choice questions

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1) The pH of a solution that contains 0.818 M acetic acid ( $K_{\rm e} = 1.77 \times 10^{-6}$ ) and 0.172 M sodium acetate is

A) 4.077 B) 5.434 C) 8.571 D) 8.370 E) 9.922

 Calculate the pH of a solution prepared by dissolving 0.37 mol of formic acid (HCO<sub>2</sub>H) and 0.23 mol of sodium formate (NaCO<sub>2</sub>H) in water sufficient to yield 1.00 L of solution. The K<sub>a</sub> of formic acid is 1.77×10<sup>-4</sup>.

A) 2.1 B) 10.46 C) 3.55 D) 2.31 E) 3.95

 Calculate the pH of a solution prepared by dissolving 0.25 mol of benzoic acid (C<sub>2</sub>H<sub>6</sub>O<sub>2</sub>H) and 0.15 of sodium benzoate (NaC<sub>2</sub>H<sub>6</sub>O<sub>2</sub>) in water sufficient to yield 1.00 L of solution. The K<sub>a</sub> of benzoic acid is 6.50×10<sup>-6</sup>.

A) 4.41 B) 2.40 C) 3.97 D) 10.08 E) 4.19

 Calculate the pH of a solution prepared by dissolving 0.15 mol of benzoic acid (HBz) and 0.30 mol of sodium benzoate in water sufficient to yield 1.00L of solution. The K<sub>s</sub> of benzoic acid is 6.50×10<sup>-6</sup>.

A) 2.52 B) 3.89 C) 4.49 D) 10.16 E) 4.20

5) Calculate the pH of a solution prepared by dissolving 0.750 mol of NH<sub>2</sub> and 0.250 mol of NH<sub>4</sub>Cl in water sufficient to yield 1.00 L of solution. The K<sub>3</sub> of ammonia is 1.77×10<sup>-5</sup>.
A) 5.22 B) 4.27 C) 9.73 D) 8.78 E) 0.89

6) The pH of a solution prepared by dissolving 0.350 mol of solid methylamine hydrochloride (CH<sub>2</sub>NH<sub>2</sub>CI) in 1.00L of 1.10M methylamine (CH<sub>2</sub>NH<sub>2</sub>) is \_\_\_\_\_\_. The K<sub>6</sub> for methylamine is  $4.4 \times 10^{-4}$ . (Assume the final volume is 1.00 L.)

A) 1.86 B) 2.66 C) 10.15 D) 11.14 E) 10.64

#### Buffers

7) What does a buffer do?

a) Keeps the pH of a solution constant

# **Practice Multiple Choice For Chem Aqueous Equilibria**

**Lingsheng Yao** 

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