

## Chapter 7 Chemical Reactions

**Section 7.1 Describing Reactions**

(pages 192–198)

*This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.*

**Reading Strategy (page 192)**

**Monitoring Your Understanding** Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

**Chemical Equations (pages 192–193)**

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. true
- Circle the letter of each sentence that is a correct interpretation of the chemical equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ .
  - Carbon and oxygen react and form carbon monoxide.
  - ☒ Carbon and oxygen react and form carbon dioxide.
  - Carbon dioxide yields carbon and oxygen.
  - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. true
- Circle the letter of the correct answer. According to the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
  - 1
  - 7
  - ☒ 14
  - 28
- In the reaction represented by the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

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**Richard Bailey**



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